Department of Pharmaceutical Chemistry, Telangana University
5yrs Integrated M.Sc Chemistry (Pharmaceutical Chemistry)
Syllabus for I and II Semesters

The following syllabus for **5yrs Integrated M.Sc Chemistry (Pharmaceutical Chemistry)** was unanimously approved by board of studies in its meeting held at Department of Pharmaceutical Chemistry, Telangana University, Nizamabad on 13th July, 2015.

Yr	Sem	Paper Code	Subject	hrs/ week	Marks		Duration of exam in hrs		Credits
					IE*	UE**	IE*	UE**	
I	**I**	**IPCH-1.1T**	English-I	4	30	70	1	3	4
		IPCH-1.2T	Second Language (Telugu/ Hindi)-I	4	30	70	1	3	4
		IPCH-1.3T	Atomic structure, Bonding, General Organic Chemistry & Aliphatic Hydrocarbons	4	30	70	1	3	4
		IPCH-1.4T	Basics of Pharmaceutical Sciences	4	30	70	1	3	4
		IPCH-1.5T	Microbiology and Immunology	4	30	70	1	3	4
		IPCH-1.3Lab	Inorganic and Organic Chemistry Lab	3	15	35	3	3	2
		IPCH-1.4Lab	Pharmaceutical Sciences Lab	3	15	35	3	3	2
		IPCH-1.5Lab	Microbiology and Immunology Lab	3	15	35	3	3	2
				21* **	**195**	**455**			**26**
I	**II**	**IPCH-2.1T**	English-II	4	30	70	1	3	4
		IPCH-2.2T	Second Language (Telugu/ Hindi)-II	4	30	70	1	3	4
		IPCH-2.3T	Chemical Energetics, States of Matter & Functional Group Organic Chemistry	4	30	70	1	3	4
		IPCH-2.4T	Anatomy and Physiology	4	30	70	1	3	4
		IPCH-2.5T	Biochemistry	4	30	70	1	3	4
		IPCH-2.3Lab	Physical and Organic Chemistry Lab	3	15	35	3	3	2
		IPCH-2.4Lab	Anatomy and Physiology Lab	3	15	35	3	3	2
		IPCH-2.5Lab	Biochemistry Lab	3	15	35	3	3	2
				21* **	**195**	**455**			**26**
			Total Marks (I and II Sem)		**1300**		**Total Credits**		**52**

*** English and SL (each 4 hrs/week) work load not included here; **University Examination; *Internal Examination

Members Present in the board of studies meeting held:

S.No	Name	Designation
1	Dr. V. Chandrasekhar, Dept. of Pharmaceutical Chemistry, Telangana University	Chairman
2	Prof. Ch. Krishna Reddy, Dept. of Chemistry, Osmania University	Member
3	Prof. V. Ravinder, Dept. of Chemistry, Kakatiya University	Member
4	Prof. C. Veeresham, UCPSc, Kakatiya University	Member
5	Dr. B. Prabhasankar, Leads Pharma Pvt. Ltd, Hyderabad	Member
6	Prof. A. K. D. Bhavani, Dept. of Chemistry, Osmania University	Member
7	Prof. Naseem, Dept. of Pharmaceutical Chemistry, Telangana University	Member
8	Prof. Devadas, Dept. of Chemistry, Osmania University	Member
9	Prof. Veerasomaiah, Dept. of Chemistry, Osmania University	Member
10	Dr. P. Jalapathi, Dept. of Chemistry, Osmania University	Member
11	Dr. B. Shireesha, Dept. of Pharmaceutical Chemistry, Telangana University	Member
12	Dr. A. Hari Padma Sree, Dept. of Chemistry, Koti Women's Coll., Osmania University	Member
13	Dr. M. Satyanarayana, Dept. of Pharmaceutical Chemistry, Telangana University	Member

IPCH-1.3T: ATOMIC STRUCTURE, BONDING, GENERAL ORGANIC CHEMISTRY & ALIPHATIC HYDROCARBONS
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Atomic Structure
Review of Bohr's theory and its limitations, dual behavior of matter and radiation, de-Broglie's relation, Heisenberg Uncertainty principle. Hydrogen atom spectra, Need of a new approach to Atomic structure.

What is Quantum mechanics? Time independent Schrodinger equation and meaning of various terms in it, Significance of ψ and ψ^2, Schrödinger equation for hydrogen atom, Radial and angular parts of the hydrogenic wave functions (atomic orbitals) and their variations for 1s, 2s, 2p, 3s, 3p and 3d orbitals (Only graphical representation), Radial and angular nodes and their significance, Radial distribution functions and the concept of the most probable distance with special reference to 1s and 2s atomic orbitals, Significance of quantum numbers, orbital angular momentum and quantum numbers m_l and m_s, Shapes of s, p and d atomic orbitals, nodal planes, Discovery of spin, spin quantum number (s) and magnetic spin quantum number (m_s)

Rules for filling electrons in various orbitals, Electronic configurations of the atoms, Stability of half-filled and completely filled orbitals, concept of exchange energy, Relative energies of atomic orbitals, Anomalous electronic configuration.

Unit-II: Chemical Bonding and Molecular Structure-I
Ionic Bonding: General characteristics of ionic bonding. Energy considerations in ionic bonding, lattice energy and solvation energy and their importance in the context of stability and solubility of ionic compounds, Statement of Born-Landé equation for calculation of lattice energy, Born -Haber cycle and its applications, polarizing power and polarizability. Fajan's rules, ionic character in covalent compounds, bond moment, dipole moment and percentage ionic character.

Unit-III: Chemical Bonding and Molecular Structure-II
Covalent bonding: Valence bond theory, Shapes of some inorganic molecules and ions on the basis of VSEPR and hybridization with suitable examples of linear, trigonal planar, square planar, tetrahedral, trigonal bipyramidal and octahedral arrangements

Concept of resonance and resonating structures in various inorganic and organic compounds

MO Approach: Rules for the LCAO method, bonding and antibonding MOs and their characteristics for s-s, s-p and p-p combinations of atomic orbitals, nonbonding combination of orbitals, MO treatment of homonuclear diatomic molecules of 1st and 2nd periods (including idea of s-p mixing) and heteronuclear diatomic molecules such as CO, NO and NO^+. Comparison of VB and MO approaches.

Unit-IV: Fundamentals of Organic Chemistry
Brief review of structural theory of organic chemistry, hybridization, bond length, bond angle, bond energy; inductive effect, electromeric effect, resonance, hyperconjugation, and their application in the

analysis of strength of organic acids, bases and stability of organic compounds; structure, shape and reactivity of nucleophiles, electrophiles and free radicals; cleavage of bonds-homolysis and heterolysis

Types of organic reactions: Addition reactions-electrophilic, nucleophilic and free radical; Substitution reactions-electrophilic, nucleophilic and free radical; Elimination and Rearrangement reactions

Unit-V: Aliphatic Hydrocarbons

Functional group approach for the following reactions (preparations & reactions) to be studied in context to their structure

Alkanes: Preparation: Catalytic hydrogenation, Wurtz reaction, Kolbe's synthesis, from Grignard reagent. Reactions: Free radical Substitution: Halogenation, Synthesis of cycloalkanes and different kinds of strains in cycloalkanes

Alkenes: Preparation: Elimination reactions: Dehydration of alkenes and dehydrohalogenation of alkyl halides (Saytzeff's rule); cis alkenes (Partial catalytic hydrogenation) and trans alkenes (Birch reduction). Reactions: cis-addition (alk. KMnO4) and trans-addition (bromine), Addition of HX (Markownikoff's and anti-Markownikoff's addition), Hydration, Ozonolysis, oxymecuration-demercuration, Hydroboration-oxidation

Alkynes: Preparation: Acetylene from CaC_2 and conversion into higher alkynes; by dehalogenation of tetra halides and dehydrohalogenation of vicinal-dihalides. Reactions: formation of metal acetylides, addition of bromine and alkaline $KMnO_4$, ozonolysis and oxidation with hot alk. $KMnO_4$.

Reference Books

1. J. D. Lee: A new Concise Inorganic Chemistry, E L. B. S.
2. F. A. Cotton & G. Wilkinson: Basic Inorganic Chemistry, John Wiley.
3. Douglas, McDaniel and Alexader: Concepts and Models in Inorganic Chemistry, John Wiley.
4. James E. Huheey, Ellen Keiter and Richard Keiter: Inorganic Chemistry: Principles of Structure and Reactivity, Pearson Publication.
5. T. W. Graham Solomon: Organic Chemistry, John Wiley and Sons.
6. Peter Sykes: A Guide Book to Mechanism in Organic Chemistry, Orient Longman.
7. E. L. Eliel: Stereochemistry of Carbon Compounds, Tata McGraw Hill.
8. I. L. Finar: Organic Chemistry (Vol. I & II), E. L. B. S.
9. R. T. Morrison & R. N. Boyd: Organic Chemistry, Prentice Hall.
10. Arun Bahl and B. S. Bahl: Advanced Organic Chemistry, S. Chand.

IPCH-1.4T: BASICS OF PHARMACEUTICAL SCIENCES
(Units: 5, Hours: 60, Credits: 4)

Basic introduction to the following topics in each unit

Unit-I: Scope of Pharmaceutical Sciences

Origin, development and scope of chemical sciences, pharmaceutical sciences, introduction to the fundamentals of pharmaceutical sciences, basic introduction, definitions, significance and interdisciplinary relevance of following in pharmaceutical sciences: Inorganic chemistry, organic chemistry, physical chemistry, analytical chemistry, biochemistry, statistics, microbiology, immunology, anatomy and physiology, pharmacology, pharmacognosy, phytochemistry, pharmaceutics, medicinal chemistry, pharmaceutical chemistry, biopharmaceutics and pharmacokinetics, clinical pharmacy/ pharmacy practice, food technology, cosmetic technology, toxicology, informatics.

Unit-II: Drugs and Pharmaceuticals

History and evolution of drugs, Ayurveda, Homeopathy, Unani, Siddha and Allopathy systems of medicine, advantages and limitations of various systems of medicine, Definition and nomenclature of pharmaceuticals, generic and trade names, source and classification of drugs based on therapeutic action, General mechanism and factors affecting drug action, pharmacodynamics, pharmacokinetics (ADME), bioavailability, agonism, antagonism, synergism, adverse reactions, Drug dependences, drug abuse, drug tolerance, hypersensitivity/ anaphylaxis.

Unit-III: Drug Discovery

Introduction to stages of drug discovery and development, lead, pharmacophore, serendipity, specific, nonspecific drugs, drug targets, activity and affinity of drugs, measurement and units of drug activity, potency of drugs, effect of concentration on activity, factors effecting drug dosage, physico-chemical factors affecting drug action, natural, semisynthetic and synthetic drugs with examples.

Unit-IV: Drug Formulations

Introduction to dosage forms, their classification with examples, routes of administration, advantages and limitations of specific dosage forms, Ayurveda, Homeopathy, Unani and Siddha formulations, Introduction, classification and uses of following additives in formulation of different dosage forms: preservatives, antioxidants, surfactants, hydrocolloids, emulsifying agents, suspending agents, diluents, binders, lubricants, and organoleptic additives.

Unit-V: Quality Control of Pharmaceuticals

Introduction to Pharmacopoeias-IP, BP, USP & International Pharmacopoeia, National Formularies and Extra Pharmacopoeia; Typical parts of a monograph of Indian pharmacopoeia with examples, quality control and quality assurance, introduction to GLP, GMP, Laboratory Accreditation, quality estimation of aspirin, acetaminophen, isoniazid, ascorbic acid, codeine phosphate, Chloride in Ringers lactate, ethambutol.

Reference Books

1. Allen, Loyd V., Jr and Howard C. Ansel: Ansel's Pharmaceutical Dosage Forms and Drug Delivery Systems, 9th Edition, Wolters Kluwer Health publisher.
2. Gareth Thomas: Med Chem: An Introduction (Kindle Edition), Wiley India Pvt Ltd; 2nd Edn.
3. Brahmankar, D.M and Sunil B. Jaiswal: Biopharmaceutics and Pharmacokinteics.
4. Mehta, R. M: Dispensing and General Pharmacy.
5. Tripathi, K. D: Essentials of Medicinal Pharmacology; 6th Edition.
6. Indian Pharmacopoeia.
7. Kokate: Pharmacognosy.
8. John H Block; John M Beale: Wilson and Gisvold's textbook of organic medicinal and pharmaceutical chemistry, 11th edn.

IPCH-1.5T: MICROBIOLOGY AND IMMUNOLOGY
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Microbial World and Microscopy

History, Scope, Future prospective of Microbiology and its application in Pharmaceutical sciences, Cell as a basic unit of life- Ultra structure of Prokaryotic cell, Ultra structure of Eukaryotic cell, Outline of classification of microorganisms, Characteristic features and economical importance of Algae, Fungi, Actinomycetes; Viruses – Structure, Classification, Medical importance; Principles of Microscopy - Light microscopy, Preparation and staining of specimens: Fixation, Dyes, Staining methods- Simple staining, Differential staining: Gram's staining, Acid-fast staining, Staining specific structures: Endospore, Capsule and Flagella staining, Fluorescent microscope, Phase contrast microscope, Confocal microscope and Electron microscope.

Unit-II: Microbial Nutrition, Culture and Growth

Culture Media, Classification and different types of media, Culture of microorganisms-Pure culture techniques and Aseptic techniques, Maintenance and preservation of Microorganisms, Microbial growth requirements, Nutritional groups of microorganisms, Bacterial growth: Bacterial cell division-Binary fission, Growth of bacterial populations: Growth curve and Growth kinetics, Measurement of growth: Direct and Indirect methods, Influence of environmental factors on Microbial growth-Temperature, pH, Osmolality and Oxygen, Cultivation of Viruses, Life cycle of Viruses: Lytic and Lysogeny

Unit-III: Microbial Growth Control

Concepts of sterilization, Physical methods of sterilization-Heat, Radiation, Filtration, Chemical agents of sterilization-Phenols and Phenolics, Alcohols, Halogens, Aldehydes, Organic acids, Surface active agents, Oxidizing agents, Gaseous agents, Evaluation of the efficiency of sterilization methods, sterility indicators, factors affecting the microbial spoilage, sources, types of microbial contaminants of pharmaceutical products, evaluation of microbial stability of formulations, Sterility testing of parenteral formulations - solids, liquids, ophthalmic and other sterile products according to the I.P., sterility testing of sterile surgical devices, Disinfectants, Sanitizers, Antiseptics and antimicrobial agents (antibacterial agents, antiviral agents, antibiotics) Effect of antimicrobial agents on Microbial growth, Measurement of antimicrobial activity (MIC), Preservation of pharmaceutical products using antimicrobial agents.

Unit-IV: Immunology

History and Concepts of immunity, Cells and Organs of the immune system, Properties of Antigens-Epitope, Hapten, Adjuvants, Structure and different classes of immunoglobulins, Cell and humoral mediated immunity, Hypersensitivity, types of allergic reactions, Antigen - Antibody reactions: Agglutination, Precipitation, Neutralization, Complement fixation, ELISA, Radioimmuno assay, Immunoblotting; Autoimmune disorders – Hashimoto's Thyroiditis, Autoimmune anemia, Insulin dependent Diabetes Mellitus, Systemic lupus erythematosus, Rheumatoid arthritis.

Unit-V: Epidemiology

Terminology of Epidemics-Sporadic, endemic, hyperendemic, outbreak, epidemic and pandemic; Tools of epidemics-Morbidity rate, prevalence rate, mortality; Recognition of an infectious disease in a population, Recognition of an epidemic; the infectious disease cycle-pathogen and reservoir; Modes of transmission of pathogen; Airborne bacterial diseases- Tuberculosis, Diphtheria, viral diseases-Chicken pox and Shingles , Flu; Water and foodborne bacterial diseases- Typhoid, Cholera, viral diseases-Polio, Hepatitis; Diseases caused by direct contact: bacterial-Leprosy, Gas gangrene, viral-AIDS, Coldsores: Zoonotic diseases: bacterial-Anthrax, Brucellosis, viral- Rabies, Ebola; Arthopod borne diseases: bacterial-Plague, Typhus fever, viral-Encephalitis, Yellow fever; Control of epidemics Vaccines and immunization

Reference Books

1. Thomas D. Brock: Biology of Microorganisms.
2. Prescott's Microbiology.
3. Gerard J. Tortora and Berdell R. Funke: Microbiology, An introduction.
4. Michael J. Pelczar: Microbiology.
5. Ananthanarayan and Paniker's Textbook of Microbiology.
6. Dr. C. B. Powar and Dr. H. F. Daginawala: General Microbiology.
7. Thomas J. Kindt, Barbara A. Osborne, Richard A. Goldsby: Kuby Immunology.
8. Peter J. Delves, Seamus J. Martin, Dennis R. Burton, Ivan M. Roitt: Essential Immunology.
9. Fahim Halim Khan: Elements of Immunology.
10. Rastogi S. C: Elements of Immonology.

IPCH-1.3LAB: INORGANIC AND ORGANIC CHEMISTRY LAB
(Credits: 2, 3hrs/week)

Section-A: Inorganic Chemistry-Volumetric Analysis

1. Estimation of sodium carbonate and sodium hydrogen carbonate present in a mixture.
2. Estimation of oxalic acid by titrating it with $KMnO_4$.
3. Estimation of water of crystallization in Mohr's salt by titrating with $KMnO_4$.
4. Estimation of Fe (II) ions by titrating it with $K_2Cr_2O_7$ using internal indicator.
5. Estimation of Cu (II) ions iodometrically using $Na_2S_2O_3$.

Section-B: Organic Chemistry

1. Solubility test of organic compounds in different test solvents & solutions
2. Detection of extra elements (N, S, Cl, Br, I) in organic compounds (containing up to two extra elements)
3. Separation of mixtures by Thin Layer Chromatography: Measure the R_f value in each case
   Identify and separate the components of a given mixture of 2 amino acids (glycine, aspartic acid, glutamic acid, tyrosine or any other amino acid) by paper chromatography

Reference Books

1. Vogel's Qualitative Inorganic Analysis, A.I. Vogel, Prentice Hall, 7th Edition.
2. Vogel's Quantitative Chemical Analysis, A.I. Vogel, Prentice Hall, 6th Edition.
3. Textbook of Practical Organic Chemistry, A.I. Vogel , Prentice Hall, 5th edition
4. Practical Organic Chemistry, F. G. Mann. & B. C. Saunders, Orient Longman, 1960

IPCH-1.4LAB: PHARMACEUTICAL SCIENCES LAB
(Credits: 2, 3hrs/week)

A brief introduction to laboratory safety, types and quality of glassware, chemicals, errors, classification of errors, accuracy and precision, minimization of errors, weighing methods, sensitivity, calibration and standard operating procedures (SOP's) of instruments (glassware, balance) with examples, normality, molality, molarity, percentage calculations, % W/V, V/V & W/W, ppm, ppb, inter conversions, isotonicity and pH of solutions with examples,

1. Preparation and dilution of solutions
2. Weighing of solid, liquid substances
3. Weight variation analysis of tablets
4. Preparation and labelling of moisturizing lotion
5. Effect of suspending agent on stability of suspensions
6. Calibration of instruments: burette, pipette, balance, thermometer, volumetric flask.
7. Preparation of neutral, acidic and basic phosphate buffers
8. Generation of drug libraries using softwares
9. Identification of pharmacophoric features and lead molecules from various drugs
10. Effect of solubility and concentration on activity of compounds
11. Estimation and comparison of potency of any two inorganic antacids available in market
12. Determination of IC_{50} from activity data of compounds
13. Extraction of caffeine from tea leaves
14. Preparation and application of Triphala churna powder
15. Preparation of herbal tea for reduction of fat content
16. Monograph analysis of paracetamol tablets

References Books

1. Allen, Loyd V., Jr, Remington: the Science and Practice of Pharmacy, 22nd Edn, Pharmaceutical Press.
2. Sidney James Carter, Cooper and Gunn's Tutorial Pharmacy, 6th Edn, Pitman Medical Publishers.
3. J. Mendham and R. C. Denney: Vogel's Quantitation Chemical Analysis, 6th Edn, Prentice Hall
4. A. H. Beckett and J. B. Stenlake: Practical Pharmaceutical Analysis, Part 1, 4th Edn, Continuum International-Athlone Publishers..

IPCH-1.5LAB: MICROBIOLOGY AND IMMUNOLOGY LAB
(Credits: 2, 3hrs/week)

1. Concepts of microscopy
2. Sterilization techniques
3. Microbiological media preparation-Nutrient Agar, Nutrient Broth
4. Isolation of pure cultures - Spread plate method, Streak plate method, Pour plate method
5. Simple staining of Bacteria
6. Gram's staining of Bacteria
7. Turbidometric estimation of bacterial growth
8. Effect of temperature on bacterial growth
9. Effect of pH on bacterial growth
10. Isolation of Fungi
11. Microscopic observation of Fungi by fungal mounting
12. Isolation of Algae
13. Rideal-Walker test
14. Widal test
15. Microscopic observation of permanent slides of Algae and Fungi
16. Study of symptoms of viral diseases

Reference Books

1. Gopal Reddy et al: Laboratory Experiments in Microbiology.
2. Arthi Nigam and Archana Ayyagari: Lab manual in Biochemistry, Immunology and Biotechnology.

IPCH-2.3T: CHEMICAL ENERGETICS, STATES OF MATTER AND FUNCTIONAL GROUP ORGANIC CHEMISTRY
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Thermodynamics-I

Thermodynamic system and surroundings, thermodynamic processes. Extensive and Intensive properties, |State and path function, Internal energy, Work and Heat. Internal energy a state function and exact differential. Work and Heat path functions and inexact differentials.

First Law of thermodynamics-Statement and Mathematical expression-its various forms.

Reversible and Irreversible process. Pressure-Volume work. Derivation of expression for maximum work done in reversible isothermal expansion of an Ideal gas, work done in irreversible isothermal expansion of Ideal gas-Problems. Derivation of expression for work done in reversible adiabatic expansion of an Ideal gas-Comparison between reversible isothermal and adiabatic expansion of an ideal gas, PV^{γ}=Constant for adiabatic changes. P,V curves for isothermal and adiabatic process.

Enthalpy-Heat of reaction at constant volume and constant pressure. Heat capacity-Heat capacities at constant volume and constant pressure. Derive Cp-Cv=R.

Variation of Heat of reaction with temperature-Kirchoff's equation and problems.

Limitations of First law of thermodynamics and need for Second law. Statement of Second law of thermodynamics-its various statements. Conversion of heat into work- Carnot cycle and efficiency of heat engine-Carnot Theorem-problems.

Concept of Entropy-entropy as a state function-Entropy changes in i) reversible process ii) Irreversible process iii) phase changes iv) Reversible isothermal expansion.

Free energy Gibbs function and Helmholtz function. Criteria for spontaneity in terms of Gibbs free energy change. Derivation of Gibbs-Helmholtz equation and its significance.

Unit-II: Gaseous and Solid State Chemistry

Gaseous State: Deviation of real gases from ideal behavior. van der Waals equation of state. Critical phenomenon. PV isotherms of real gases, continuity of state. Andrew's isotherms of CO_2. The van der Waal's equation and critical state. Derivation of relationship between critical constants and van der Waal's constants. The law of corresponding states, reduced equation of states. Joule Thomson effect and inversion temperature of a gas. Liquifaction of gases: i) Linde's method based on Joule Thomson effect ii) Claude's method based on adiabatic expansion of a gas.

Solid State: Laws of Crystallography – (i) Law of Constancy of interfacial angles (ii) Law of Symmetry, Symmetry elements in crystals (iii) Law of rationality of indices. Definition of space lattice, unit cell. Bravis Lattices and seven crystal systems (a brief review). X-ray diffraction by crystals; Derivation of Bragg's equation and its application in the determination of structure of NaCl, KCl & CsCl (Bragg's method and Powder method). Frenkel and Shottky defects. Band structure of solids.

Unit-III: Benzene, Alkyl and Aryl Halides

Benzene: Preparations: from phenol, by decarboxylation, from acetylene, from benzene sulphonic acid. Reactions (Case benzene): Orientation in aromatic substitution reactions (ring activation and ring deactivation) Electrophilic substitution (with mechanism) reactions nitration, halogenation and sulphonation, Friedel-Craft's reaction (alkylation and acylation), Side chain oxidation of alkyl benzenes, Aromaticity: Benzenoids and Hückel's rule

Alkyl Halides: Types of Nucleophilic Substitution (S_N1, S_N2 and S_Ni) reactions

Preparation: from alkenes and alcohols. Reactions: hydrolysis, nitrite & nitro formation, nitrile & isonitrile formation. Williamson's ether synthesis: Elimination Vs substitution.

Aryl Halides: Preparation: (Chloro, bromo and iodo-benzene case): from phenol, Sandmeyer & Gattermann reactions. Reactions (Chlorobenzene): Aromatic nucleophilic substitution (replacement by -OH group) and effect of nitro substituent. Benzyne Mechanism: KNH_2/NH_3 (or $NaNH_2/NH_3$).

Reactivity and Relative strength of C-Halogen bond in alkyl, allyl, benzyl, vinyl and aryl halides.

Unit-IV: Alcohols, Phenols and Ethers

Alcohols: Preparation: Preparation of 1°, 2° and 3° alcohols: using Grignard reagent, Ester hydrolysis, Reduction of aldehydes, ketones, carboxylic acid and esters. Reactions: With sodium, HX (Lucas test), esterification, oxidation (with PCC, alk. $KMnO_4$, acidic dichromate, conc. HNO_3), Oppeneauer oxidation Diols: (Upto 6 Carbons) oxidation of diols. Pinacol-Pinacolone rearrangement

Phenols: Preparation: (Phenol case) Cumene hydroperoxide method, from diazonium salts. Reactions: Electrophilic substitution: Nitration, halogenation and sulphonation. Reimer-Tiemann Reaction, Gattermann-Koch Reaction, Houben-Hoesch Condensation, Schotten-Baumann Reaction.

Ethers (aliphatic and aromatic): Introduction, synthesis and reactions of ethers

Unit-V: Aldehydes and ketones

(Formaldehye, acetaldehyde, acetone and benzaldehyde)

Preparation: from acid chlorides and from nitriles. Reactions: Reaction with HCN, ROH, $NaHSO_3$, NH_2-G derivatives. Aldol Condensation, Cannizzaro's reaction, Wittig reaction, Benzoin condensation, Clemensen reduction and Wolff Kishner reduction, Meerwein-Pondorff Verley reduction, Analysis of aldehydes and ketones: haloform test, 2,4-DNP test, Tollens and Fehling test

Reference Books

1. T. W. Graham Solomons: Organic Chemistry, John Wiley and Sons.
2. Peter Sykes: A Guide Book to Mechanism in Organic Chemistry, Orient Longman.
3. I.L. Finar: Organic Chemistry (Vol. I & II), E. L. B. S.
4. R. T. Morrison & R. N. Boyd: Organic Chemistry, Prentice Hall.
5. G. W. Castellan: Physical Chemistry 4th Edn. Narosa (2004).
6. Rajaram: Thermodynamics.
7. Puri and Sharma: Text Book of Physical Chemistry.
8. Puri, Sharma and Pathania: Text Book of Physical Chemistry.
9. Gurdeep Raj: Advanced Physical Chemistry.

IPCH-2.4T: ANATOMY AND PHYSIOLOGY
(Units: 5, Hours: 60, Credits: 4)

Unit-I: **Blood, Lymphatic and Skeletal system**

Blood: Composition and functions of blood, RBC, WBC, platelets, homeostasis, blood groups, mechanism of clotting, Introduction to disorders of blood; **Lymphatic System:** Composition, formation and circulation of lymph, Spleen and its functions; **Skeletal System:** Structure and function of skeleton, Articulation and movement, Disorders of bones and joints.

Unit-II: GIT, Respiratory System and Sensory Organs

Gastrointestinal tract: Structure of the gastrointestinal tract, functions of its different parts including those of liver, pancreas and gall bladder, various gastrointestinal structures and their role in the digestion and absorption of food; **Respiratory System:** Structure of respiratory organs, functions of respiration mechanism and regulation of respiration, respiratory volumes and vital capacity; **Sensory Organs:** Structure and physiology of eye (vision), ear (hearing), taste buds, nose (smell) and skin.

Unit-III: **Nervous System**

Autonomic Nervous System: Physiology and functions of the autonomic nervous system, Mechanism of neurohumoral transmission in ANS; **Central Nervous System:** Functions of different parts of brain and spinal cord, Neurohumoral transmission in the central nervous system, reflex action, electroencephalogram, specialized functions of the brain, cranial nerves and their functions

Unit-IV: **Urinary and Cardiovascular Systems**

Urinary System: Structure and functions of the kidney and urinary tract, physiology of urine formation and acid base balance, brief introduction to disorders of kidney; **Cardio Vascular System:** Anatomy of heart and blood vessels, physiology of blood circulation, cardiac cycle, conducting system of heart, heart sound, electrocardiogram, blood pressure and its regulation.

Unit-V: Endocrine and Reproductive Systems

Endocrine Glands: Basic anatomy and physiology of pituitary, thyroid, parathyroid, adrenal glands and pancreas, local hormones, brief introduction to disorders of various endocrine glands; **Reproductive System:** Structure and functions of male and female reproductive system, sex hormones, physiology of menstrual cycle, and various stages of pregnancy and parturition.

Reference Books

1. Vander, Sherman and Luciano: Human Physiology, the Mechanism of Body Function.
2. John B. West: Best and Taylor's Physiological Basis of Medical Practice.
3. Gerard J. Tortora & Bryan Derikson: Principles of Anatomy and Physiology.
4. Arthur C. Guyton and John E. Hall: Text Book of Medical Physiology, Elsevier India.
5. Anne Waugh Allison Grant Ross and Wilson, Churchill Livingstone: Anatomy and Physiology in Health and Illness, Elsevier Publishers.

IPCH-2.5T: BIOCHEMISTRY
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Carbohydrates

Classification, general properties and biological importance of carbohydrates; Monosaccharides: Fischer and Haworth projections (open chain and cyclic structure) and stereoisomerism of Glucose and Fructose; Determination of configuration of monosaccharides, absolute configuration of Glucose and Fructose, Mutarotation, Disaccharides: Glycosidic bond, Structure determination of Sucrose and structures of cellobiose, maltose and lactose; Polysaccharides: Structural polysaccharides-Cellulose, Chitin and Peptidoglycan, Storage polysaccharides: Starch, Glycogen and Inulin, Glycosamino glycanes, artificial sugars, Metabolism of carbohydrates: glycolysis, TCA, ETC, Glycogenesis, glycogenolysis, gluconeogenesis Hexose monophosphate shunt [HMP shunt]; Diseases associated with carbohydrate metabolism.

Unit-II: Amino Acids and Proteins

Classification and biological importance of proteins, Amino acids: Structures, isomerism, properties and classification; Structure of proteins, Ramachandran plot; Synthesis and sequencing of proteins; Protein databases; Denaturation & its effect on biological activity; Metabolism of amino acids: Transamination & deamination, decarboxylation, Urea cycle; Diseases associated with various proteins, Proteins as biomarkers and Drug targets, Protein extraction and its applications in food and drug industry.

Unit-III: Fatty Acids and Lipids

Classification and biological importance of lipids; Fatty acids: Structures, isomerism, properties and classification, Fats and oils, waxes; Triacyl glycerols, Phospholipids and Spingo lipids; Sterols; Steroids: adreno corticoids, testosterone, progesterone, estrogens; formation & biochemical significance of bile acids, bile salts, ketone bodies, prostaglandins, Biological membranes, Metabolism of fatty acids: beta oxidation; Diseases associated with lipid metabolism.

Unit-IV: Nucleic Acids

History and discovery of DNA and RNA; Chemical structure of DNA and RNA: Purine & pyrimidine bases, sugars, & phosphoric acid. Nucleosides & nucleotides, Watson-Crick double helix model of DNA, Conformation of DNA- primary, secondary, tertiary level; formation of nucleosome; structure of chromosome, karyotyping; alternate forms of DNA; DNA supercoiling, types of RNA; denaturation and renaturation of DNA; Synthesis and sequencing of DNA; Nucleic acid databases, replication of DNA, transcription, genetic code, translation, Mutations, Nucleic acids as drug targets

Unit-V: Enzymes

History, nomenclature and classification of enzymes; Properties of enzymes, role of co-factors and coenzymes; Mechanism of action of enzymes; Kinetics of enzymes, double reciprocal plot, Km and

Vmax; factors effecting enzymatic action; Enzyme inhibition. Enzymes as targets of drugs, enzymes from natural sources and their application

Reference Books

1. Donald Voet and Judith G. Voet: Biochemistry.
2. Albert L. Lehninger: Biochemistry.
3. Reginald H. Garret and Charles M. Grisham: Biochemistry.
4. Stryer: Biochemistry.
5. U. Satyanarayana: Biochemistry.
6. Karp: Cell and Molecular Biology.
7. Cooper: Cell and Molecular Biology
8. Rasthogi: Cell and Molecular Biology
9. Snustad Simmonds: Genetics.
10. Brooker: Genetics.

IPCH-2.3LAB: PHYSICAL AND ORGANIC CHEMISTRY LAB
(Credits: 2, 3hrs/week)

Section A: Physical Chemistry

Thermochemistry

1. Determination of heat capacity of calorimeter for different volumes
2. Determination of enthalpy of neutralization of hydrochloric acid with sodium hydroxide
3. Determination of enthalpy of neutralization of acetic acid with ammonium hydroxide
4. Determination of enthalpy of ionization of acetic acid
5. Determination of integral enthalpy of solution of salts (KNO_3, NH_4Cl)
6. Determination of enthalpy of hydration of copper sulphate

Section B: Organic Chemistry

Preparations:

(a) Bromination of Phenol/Aniline
(b) Benzoylation of amines/phenols
(c) Oxime and 2,4-dinitrophenylhydrazone of aldehyde/ketone

The following methods are to be applied in the preparation of the above mentioned compounds:
Purification: by crystallization/distillation
Purity check: by determination of melting point/boiling points and by TLC
Calculations: quantitative yields

Reference Books

1. A.I. Vogel: Textbook of Practical Organic Chemistry, 5th edition, Prentice-Hall.
2. F. G. Mann & B. C. Saunders: Practical Organic Chemistry, Orient Longman (1960).
3. B. D. Khosla, V. C. Garg and Gulati: Senior Practical Physical Chemistry, R. Chand & Co.
4. Alexander Findlay: Practical Physical Chemistry.
5. B. P. Levitt: Findlay's Practical Physical Chemistry.
6. P. S. Sindhu: Practicals in Physical Chemistry, A Modern Approach.

IPCH-2.4 LAB: ANATOMY AND PHYSIOLOGY LAB
(Credits: 2, 3hrs/week)

1. Microscopic study of different tissues
2. Identification of bones and joints
3. Study of different systems with the help of charts and models
4. Estimation of RBC
5. Estimation of WBC
6. Estimation of haemoglobin
7. Different leukocyte count
8. Estimation of erythrocyte sedimentation rate
9. Determination of blood group
10. Determination of bleeding time
11. Determination of clotting time
12. Recording of human heart rate and pulse rate
13. Study of effect of posture and exercise on blood pressure
14. Recording human body temperature

Reference Books

1. Shukant R. Apte: Experimental Physiology.
2. Ramesh K. Goyal, Natyar M. Patel and Shailesh A. Shah: Practical Anatomy, Physiology and Biochemistry.
3. Sir John Y. Dacie and S. M. Lewis: Practical Haematology.

IPCH-2.5LAB: BIOCHEMISTRY LAB
(Credits: 2, 3hrs/week)

1. Extraction of starch from Potato and its identification
2. Qualitative analysis of sugars
3. Qualitative analysis of Amino acids
4. Qualitative analysis of Lipids
5. Quantitative estimation of sugars by Anthrone method
6. Quantitative estimation of proteins by Biuret method
7. Extraction of DNA from Bacteria/Blood
8. Estimation of DNA by DPA method
9. Estimation of RNA by Orcinol method
10. Extraction of Amylase and determination of its activity
11. Separation and identification of amino acids by chromatography
12. Extraction of Casein from milk
13. Visualization of primary, secondary, tertiary structure of various proteins in PDB
14. Determination of blood glucose levels
15. Determination of SGOT and SGPT levels

Reference Books

1. Deshpandey and Shashidhar: Lab Manual.
2. Plummer: Lab Manual.
3. Sahney: Biochemistry Lab Manual.

Department of Pharmaceutical Chemistry, Telangana University
5yrs Integrated M.Sc Chemistry (Pharmaceutical Chemistry)
Syllabus for III and IV Semesters

The following syllabus for **5yrs Integrated M.Sc Chemistry (Pharmaceutical Chemistry)** was unanimously approved by board of studies in its meeting held at Department of Chemistry, Osmania University, Hyderabad on 27th June, 2016.

Yr	Sem	Paper Code	Subject	hrs/ week	Marks		Duration of exam in hrs		Credits
					IE*	UE**	IE*	UE**	
II	**III**	**IPCH-3.1T**	English-III	4	30	70	1	3	4
		IPCH-3.2T	Environmental Science	4	30	70	1	3	4
		IPCH-3.3T	Solutions, Electrochemistry & Functional Group Organic Chemistry	4	30	70	1	3	4
		IPCH-3.4T	Pharmaceutical Analysis	4	30	70	1	3	4
		IPCH-3.5T	Cell and Molecular Biology	4	30	70	1	3	4
		IPCH-3.3Lab	Physical Chemistry and Organic Chemistry Lab	3	15	35	3	3	2
		IPCH-3.4Lab	Pharmaceutical Analysis Lab	3	15	35	3	3	2
		IPCH-3.5Lab	Cell and Molecular Biology Lab	3	15	35	3	3	2
				25* **	**195**	**455**			**26**
II	**IV**	**IPCH-4.1T**	English-IV	4	30	70	1	3	4
		IPCH-4.2T	Chemistry of Nutraceuticals and Dairy Products	4	30	70	1	3	4
		IPCH-4.3T	Coordination Chemistry, Chemical Kinetics, Dilute Solutions & Colligative Properties, Colloids & Surface Chemistry	4	30	70	1	3	4
		IPCH-4.4T	Pharmacology-I	4	30	70	1	3	4
		IPCH-4.5T	Genetic Engineering and its Applications in Pharmaceutical Sciences	4	30	70	1	3	4
		IPCH-4.3Lab	Inorganic and Physical Chemistry Lab	3	15	35	3	3	2
		IPCH-4.4Lab	Pharmacology Lab	3	15	35	3	3	2
		IPCH-4.5Lab	Genetic Engineering Lab	3	15	35	3	3	2
				25* **	**195**	**455**			**26**
			Total Marks (III and IV Sem)		**1300**		**Total Credits**		**52**

*** English work load (4 hrs/week) not included here; **University Examination; *Internal Examination

Members Present in the Board of Studies Meeting Held:

S.No	Name	Designation
1	Dr. V. Chandrasekhar, Dept. of Pharmaceutical Chemistry, Telangana University	Chairman
2	Prof. Ch. Krishna Reddy, Dept. of Chemistry, Osmania University	Member
3	Prof. V. Ravinder, Dept. of Chemistry, Kakatiya University	Member
4	Prof. C. Veeresham, UCPSc, Kakatiya University	Member
5	Prof. Naseem, Dept. of Pharmaceutical Chemistry, Telangana University	Member
6	Prof. Devadas, Dept. of Chemistry, Osmania University	Member
7	Prof. Veerasomaiah, Dept. of Chemistry, Osmania University	Member
8	Dr. P. Jalapathi, Dept. of Chemistry, Osmania University	Member
9	Dr. A. Hari Padma Sree, Dept. of Chemistry, Koti Women's Coll., Osmania University	Member
10	Dr. M. Satyanarayana, Dept. of Pharmaceutical Chemistry, Telangana University	Member

IPCH-3.2T: ENVIRONMENTAL SCIENCE
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Earth's Spheres and Environmental Pollution

Structure of earth spheres, basics of bio-geochemical cycles, types of environmental pollution

Air Pollution: Types and sources, green house effect and global warming, measures for reversion of green house effect, acid rain, ozone depletion and their consequences on environment, photochemical smog, hazards related to power plants, chemical and biological warfare systems, strategies for air pollution control

Water Pollution: Hydrological cycle, types and sources of water pollution, composition of organic and inorganic water pollutants, ground water pollution, surface water pollution (lake and river water), marine pollution, mercury pollution, eutrophication, biodegradability of detergents and pesticides-case studies, harmful effects of water pollution and control strategies

Unit-II: Water Analysis and Waste Water Treatment

Water Analysis: Water quality criteria for domestic use (BIS), analysis of temperature, color, odor, taste, turbidity, pH (physical parameters), TDS, total alkalinity, hardness, Cl^-, F^-, residual-Cl, NH_3, NO_3^-, NO_2^-, SO_4^{-2}, CO_3^{-2}, Na, K, Mg and Ca (chemical parameters), heavy metals, microbiological and pesticide residues

Water Treatment: Principles of water and waste water treatment for drinking and industrial use-(i) physical process (ii) biological process (iii) chemical processes - principles of coagulation, flocculation, softening, disinfection, electro-dialysis, reverse osmosis, demineralization, advanced oxidation process (AOP) and fluoridation

Dissolved oxygen (DO) and water aeration, chemical oxygen demand (COD), biochemical oxygen demand (BOD), aerobic and anaerobic treatment of waste water, aeration of water, activated sludge

Unit-III: Solid and Hazardous Waste Management

Solid Wastes: Soil fertility and solid wastes definition, types, sources, characteristics, and impact of solid wastes on environmental health

Solid-waste collection, segregation, transport and handling of wastes at source,

Solid waste processing technologies; chemical, physical and biological treatment for energy and other resource recovery-recycling and reuse,

Bio-remediation, composting and vermicomposting, Landfill design for sanitary and hazardous wastes and incineration, landfill bioreactors

Hazardous Wastes: Definition, categorization, sources and characteristics, hazardous waste collection, transport, treatment and disposal, legislation on management and handling of solid and hazardous wastes, disposal of medical and pharmaceutical waste.

Unit-IV: Energy Resources and Maintenance

Renewable and non-renewable energy resources, physico-chemical characteristics and energy content of coal, petroleum and natural gas, basic concept of various renewable energy resources and

utilization - solar radiation and its spectral characteristics, solar energy for chemical synthesis, photovoltaics, wind energy, geothermal energy, ocean energy, hydroelectric energy, biomass energy & biofuels, hydrogen energy, cryogenic fuels, fuel-cells, energy storage and usage pattern, role of IT (information technology) in energy management, future energy need projection in India.
Environmental protection act; air, water, wild life and forest conservation act, basic concepts of bio-diversity and biodiversity act.

Unit-V: Nuclear and Radiochemistry

Introduction, radioactive decay and equilibrium, types of nuclear reactions, nuclear fission and nuclear reactors, nuclear fusion, nuclear power plants and fuel production
Sources of radioactive pollution, biological effects, radiation detection, nuclear fallout and disasters examples, control of radiation/disposal options, basics of radiochemical separation, basics of tracer applications, radio carbon dating
Concept of environmental management and disaster reduction

Reference Books

1. Bureau of Indian Standards (BIS) IS:10500:2012, Drinking Water Specification, 2012.
1. A. K. De: Environmental Pollution.
2. Wark and Werner: Air Pollution.
3. B. K. Sharma & H. Kaur: Environmental Pollution.
4. P. K. Trivedi: Introduction to Air Pollution.
5. D. D. Tyagi and M. Mehre: A Text Book of Environmental Pollution.
6. C. S. Rao: Environmental Pollution Engineering and Control.
7. Satake and M. Midu: Chemical in the Environment.
8. G. Lunn and E. B. Sansone: Destruction of hazardous chemicals in the laboratory
9. A. E. Greenberg and A. D. Eaton: Standard Methods for Examination of Water and Waste Water, APHA, AWWA,WEF
10. D. S. Ramteke and C. A. Moghe: Manual on Water and Waste water analysis, NEERI
11. H. Kaur, Environmental Chemistry, 6th Edn, Pragathi Prakashan, Meerut, 2011
12. K. H. Mancy and W. J. Weber Jr: Analysis of Industrial Waste Water, Wiley Interescience, New York, 1971

PCH-3.3T: SOLUTIONS, ELECTROCHEMISTRY & FUNCTIONAL GROUP ORGANIC CHEMISTRY

(Units: 5, Hours: 60, Credits: 4)

Unit-I: Solutions and Phase Rule

Solutions: Liquid-liquid mixtures, ideal liquid mixtures**,** Raoult's and Henry's laws. Non-ideal systems. Azeotropes HCl-H_2O and C_2H_5OH-H_2O systems. Fractional distillation, Partially miscible liquids- Phenol-Water, Triethyl amine-Water and Nicotine-Water systems. Lower upper consolute temperatures. Effect of impurity on consolute temperature. Immiscible liquids and steam distillation.

Phase Rule: Statement and meaning of the terms-Phase, Component and degrees of freedom, Gibb's Phase rule, phase equilibria of one component system-water system. Phase equilibria of two-component system- Solid-Liquid equilibria, simple eutectic Pb-Ag system, desilverisation of lead. Solid solutions- compound with congruent melting point Mg-Zn system and incongruent melting point NaCl-H_2O system

Unit-II: Electrochemistry

Electrical transport-conduction in metals and in electrolyte solutions, Specific conductance, Equivalent conductance and molecular conductance, measurement of equivalent conductance, variation of specific conductance and equivalent conductance with dilution. Migration of ions-Relative speed of ions during electrolysis. Transport Number-definition and determination by Hittorf's method for attackable and unattackable electrodes. Kohlraush law of independent migration of ions and applications. Theory of electrolyte dissociation and its limitations, weak and strong electrolytes, Ostwald's dilution law-its uses and limitations. Debye-Huckel theory of strong electrolytes, Debye-Huckel-Onsager equation for stronge electrolytes (elementary treatment only)

Difference between electrolytic cell and Galvanic cell. Galvanic cells- Conventional cell representation of Galvanic cells, Single electrode potential, Standard electrode potential.

EMF of a cell and its measurement, reference electrodes-Standard hydrogen electrode (SHE) and Calomel electrode. Electrode reactions-Nernst equation. Reversible and Irreversible cells, Types of reversible electrodes-gas electrode, metal-metal ion electrode, metal-metal insoluble salt electrode and redox electrodes. Calculation of thermodynamic quantities of cell reactions (ΔG, ΔH and K)

Unit-III: Carboxylic acids, Amines and Their Derivatives

Carboxylic acids (aliphatic and aromatic)

Preparation and reactions including Hell-Vohlard-Zelinsky Reaction

Carboxylic acid derivatives (aliphatic):

Preparation: Acid chlorides, Anhydrides, Esters and Amides from acids and their interconversion.

Reactions: Comparative study of nucleophilicity of acyl derivatives. Reformatsky Reaction, Perkin condensation.

Amines and Diazonium Salts

Amines (Aliphatic and Aromatic):

Preparation: from alkyl halides, Gabriel's Phthalimide synthesis, Hofmann Bromamide reaction.

Reactions: Hofmann vs. Saytzeff elimination, Quaternary ammonium salts, Carbylamine test, Hinsberg test, test with HNO_2, Schotten-Baumann Reaction. Electrophilic substitution (case aniline): nitration, bromination, sulphonation.
Diazonium salts: Preparation: from aromatic amines.
Reactions: conversion to benzene, phenol, dyes.

Unit-IV: Amino Acids and Peptides

Preparation of Amino Acids: Strecker synthesis using Gabriel's phthalimide synthesis. Zwitterion, Isoelectric point and Electrophoresis.
Reactions of Amino acids: ester of –COOH group, acetylation of $–NH_2$ group, complexation with Cu^{2+} ions, ninhydrin test.
Determination of Primary structure of Peptides by degradation Edmann degradation (N-terminal) and C-terminal (thiohydantoin and with carboxypeptidase enzyme). Synthesis of simple peptides (up to dipeptides) using coupling reagents (EDC/HOBt, DCC, DIC, HATU) and discuss the advantages of various coupling reagents

Unit-V: Nitro Compounds

Classification, Tautomerism of nitroalkanes (nitro form and aci form), Preparation (nitroalkanes and alkyl nitrites), Reactivity: Halogenation using nitrous acid, Nef reaction, Michael Addition, Henry Reaction, Aromatic Nitro hydrocarbons, Preparation of Nitrobenzene from diazonium salt and direct nitration, Reactivity and reduction of nitrobenzene in different media

Reference Books

1. Mahan, B. H: University Chemistry, 3rd Ed. Narosa (1998).
2. Petrucci, R. H: General Chemistry, 5th Edn, Macmillan Publishing Co.: New York (1985).
3. Morrison, R. T. & Boyd, R. N: Organic Chemistry, Dorling Kindersley (India) Pvt. Ltd. (Pearson Education).
4. Finar, I. L: Organic Chemistry (vol.1), Dorling Kindersley (India) Pvt. Ltd. (Pearson Edn).
5. Finar, I. L: Organic Chemistry (vol.2), Dorling Kindersley (India) Pvt. Ltd. (Pearson Edn).
6. Nelson, D. L & Cox, M. M: Lehninger's Principles of Biochemistry, 7th Edn, W. H. Freeman.
7. Berg, J. M., Tymoczko, J. L. & Stryer, L: Biochemistry, W.H. Freeman (2002).
8. Prutton and Marron: Principles of Physical Chemistry.
9. Puri and Sharma: Text Book of Physical Chemistry.
10. S. K. Dogra: Physical Chemistry Through Problems.
11. Puri, Sharma and Pathania: Text Book of Physical Chemistry
12. Gurdeep Raj: Advanced Physical Chemistry

IPCH-3.4T: PHARMACEUTICAL ANALYSIS
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Fundamentals of Chemical Analysis

Introduction to analytical chemistry and chemical analysis, applications of chemical analysis, review of basic chemical measurements, units, and concentrations, types of analysis, classifications of analytical techniques based on sample and methods of analysis, factors affecting the choice of analytical methods, sampling, sampling procedures, principles of sample preparation, including digestion, precipitation, extraction (including solvent extraction, ion pair extraction, solid phase extraction), fusion, ashing, drying, hazards in sampling, accuracy, precision, calibration, errors, types of errors, propagation of errors, minimization of errors

Unit-II: Volumetric Analysis-1

Theoretical considerations and pharmaceutical applications with special reference to Indian Pharmacopoeia of the following analytical techniques

Acid-Base titrations: primary and secondary standards, Law of mass action, hydrolysis of salts, neutralization curves, and theory of indicators, choice of indicators, mixed indicator, application in assay of Benzoic acid, Boric acid, Aspirin as per IP.

Non-Aqueous titrations: Types of solvents, end point detection, application in assay of Sodium acetate, Sodium benzoate, Norfloxacin tablet as per IP

Oxidation-Reduction titrations: Theory of redox titration, oxidation-reduction curves, redox Indicators, pharmaceutical applications, preparation and standardization of redox titrants like potassium permanganate, ceric ammonium sulphate, sodium thiosulphate, titanous chloride, applications in assay of Ferrous sulfate, Ascorbic acid, Isoniazide, Hydrogen peroxide, Iodine (Iodimetry and Iodometry)

Unit-III: Volumetric Analysis-2

Theoretical considerations and pharmaceutical applications with special reference to Indian Pharmacopoeia of the following analytical techniques

Complexometric Titrations: Formation of complex and its stability, methods to increase the selectivity of EDTA titrations, titration curves, metallochrome indicators, types of EDTA titrations, application in assay of Magnesium sulfate, Lead nitrate and calcium gluconate.

Argentometric Titrations: theory of precipitation, factors affecting solubility of a precipitate, titration methods Mohr's, Volhard's, Gay lussac, and Fajan's method, indicators. Applications in assay of Potassium chloride, Sodium chloride and Ammonium chloride

Unit-IV: Chromatography

Principles and classification with relevant examples of pharmaceutical products involving separation of drugs from excipients or impurities using following chromatographic techniques

TLC: Introduction, principle, techniques, Rf value, resolution, mobile and stationary phases, applications and visualization methods

Column Chromatography: Adsorption column chromatography, Operational technique, frontal analysis and elution analysis. Factors affecting column efficiency, applications

Unit-V: Spectroscopic Techniques

Introduction to theory of electronic, atomic and molecular spectra, Fundamental laws of photometry, Beer-Lambert's Law, application and its deviation, limitation of Beer law, application of the law to single and multiple component analysis, Spectra of isolated chromophores, auxochromes, bathochromic shift, hypsochromic shift, hyperchromic and hypochromic effect, effect of solvent on absorption spectra, molecular structure, Photometer, U.V-Visible spectrophotometer-sources of U.V-Visible radiations, collimating systems, monochromators, samples cells and following detectors-Photocell, Barrier layer cell, Phototube, Diode array, applications of U.V-Visible spectroscopy in pharmacy and spectrophotometric titrations with examples from IP.

Reference Books

1. Gary D. Christian: Analytical Chemistry, 6th edition, John Wiley & sons Inc.
2. Douglas A. Skoog, Donald M. West, James F. Holler, Stanley R. Crouch: Fundamentals of Analytical Chemistry.
3. Banwell and McCash: Fundamentals of Molecular Spectroscopy, 4th edn, Tata McGraw Hill
4. Day and Underwood: Quantitative Analysis, 6th Edition, PHI, 2009.
5. David Harvey: Modern Analytical Chemistry, McGraw Hill, 2000,
6. Chatwal K, Anand: Instrumental Methods of Chemical Analysis, 5th edition.
7. G. H. Jeffery, J. Bassett, J. Mendhan, R. C. Denny: Vogel's text book of Quantitative Chemical Analysis.
8. Willard, Merit, Dean, Settle: Instrumental Methods of Analysis.
9. Y. R. Sharma: Organic Spectroscopy.
10. Silverstein, Webster: Spectrometric Identification of Organic Compounds.
11. Henry P. Talbot: An introductory Course of Quantitative Chemical Analysis with Explanatory Notes, 6th edition.
12. Indian Pharmacopoeia, BP, USP.
13. Remington: The Science and Practice of Pharmacy, 20th edition.
14. Donald Cairns: Essentials of Pharmaceutical Chemistry, third edn, Pharmaceutical press.
15. A. H. Becket and J. B. Stenlake: Practical Pharmaceutical Chemistry, 4th edn, the Athlone press.

IPCH-3.5T: CELL AND MOLECULAR BIOLOGY
(Units: 5, Hours: 60, Credits: 4)

Unit-I:

Cell organelles of Eukaryotic Cell: Dynamic structure of membrane and its function, Chemical composition of membrane-Membrane lipids, Proteins and carbohydrates; Membrane fluidity, Membrane transport-Passive and Active transport, Movement of ions through membrane-Ion Channels; Neurotransmission, Action of drugs on Synapses; Mitochondrial structure and function; Endoplasmic Reticulum; Golgi complex; Ribosomes; Peroxisomes; Extracellular matrix in Animal Cells, Interaction of Cells with Extracellular matrix, Interactions of Cells with other Cells

Unit-II:

Cell Cycle, Cell Cycles in Vivo, Regulation of Cell Cycle by cyclin dependent Kinases (CdKs); Mitosis, its stages and significance, Meiosis, its stages and significance, Gametogenesis, Mitotic and Meiotic nondisjunction and its consequences; Apoptosis, Extrinsic and Intrinsic Pathways of Apoptosis, Cancer, Properties of Cancer Cells, Different types of Cancers, Causes of Cancer, Genetics of cancer-Tumour-Suppressor genes and Oncogenes, Strategies for combating Cancer

Unit-III:

Signal Transduction-Cell signaling, Different types of Cell signaling systems, Signaling pathway; Extracellular messengers and their receptors, Secondary messengers and their role in Signal Transduction-c AMP, Inositol 1,4,5 triphosphate, Diacylglycerol and Calcium; Signal Transduction by G Protein Coupled Receptors and Receptors Protein-Tyrosine Kinases(RTKs) with suitable examples, Insulin Signaling defect in Diabetes mellitus

Unit-IV:

Gene as a unit of inheritance, Chromosome-Carriers of genes, Chromosome structure -centromere, Telomere, satelite DNA; Cytological techniques used to study chromosomes, Karyotype of Humans, Location of genes on chromosomes-Fluorescent insitu Hybridization; Separation of chromosomes-Fluorescent Activated Cell Sorter; Chromosomal aberrations: Structural aberrations- Inversion, Deletion, Duplication, Translation. Numerical abberations- Hypoploidy, Hyperploidy, Polyploidy; Diseases associated with chromosomal abnormalities (Autosomal/Sex chromosomal), Prenatal diagnosis of Chromosomal abnormalities

Unit-V:

The concept of Genome, Organization of genome in Virus &Bacteria, Genome organization and role of Histone proteins in Eukaryotes; The complexity of the genome -Denaturation & Renaturation, Reassociation Kinetics, Cot Curves, Human genome Project , Genetic variation within the human species population-1)DNA sequence variation,2) Copy number variation, 3)Structural variation; Application of Genome analysis to Medicine, Amplification of Gene-Polymerase Chain Reaction, Primer designing; Molecular diagnosis of Diseases

Reference Books

1. Snustad Simmons: Principles of Genetics.
2. Daniel L. Hartl and Elizabeth W Jones: Genetics.
3. Robert J. Brooker: Genetics.
4. Peter J. Russell: Genetics.
5. Gerald Karp and Nancy L Pruitt: Cell Molecular Biology.
6. Cooper: Cell and Molecular Biology.
7. David Baltimore and Harvey Lodish: Molecular Cell Biology.

IPCH-3.3LAB: PHYSICAL CHEMISTRY AND ORGANIC CHEMISTRY LAB
(Credits: 2, 3hrs/week)

Physical Chemistry

Critical solution temperature (CST)

1. Determination of CST of phenol-water system
2. Effect of salt on the CST in phenol-water system

Distribution

1. Distribution of acetic acid between n-butanol and water
2. Distribution of benzoic acid between toluene and water

Conductance

1. Determination of Cell constant
2. Verification of Ostwald's dilution law using a weak acid
3. Conductometric titration of strong acid Vs strong base
4. Conductometric titration of weak acid Vs strong base
5. Conductometric titration of a mixture of strong acid and weak acid Vs strong base

Organic Chemistry

Systematic identification of organic compounds possessing mono functional groups (-COOH, phenolic, aldehydic, ketonic, amide, nitro, amines), preparation & purification of one derivative and detection of extra elements (N, S, Cl, Br, I) present in the above mentioned functional group containing organic compounds.

Reference Books

1. Vogel, A. I., Tatchell, A. R., Furnis, B. S., Hannaford, A. J. & Smith, P. W. G: Textbook of Practical Organic Chemistry, Prentice-Hall, 5th edition (1996).
2. Mann, F. G. & Saunders, B. C: Practical Organic Chemistry Orient-Longman (1960).
3. Khosla, B. D.; Garg, V. C. & Gulati, A: Senior Practical Physical Chemistry, R. Chand & Co.
4. Ahluwalia, V.K. & Aggarwal, R: Comprehensive Practical Organic Chemistry.
5. B. D. Khosla, V. C. Garg and Gulati: Senior Practical Physical Chemistry, R. Chand & Co.
6. Alexander Findlay: Practical Physical Chemistry.
7. B. P. Levitt: Findlay's Practical Physical Chemistry.
8. P. S. Sindhu: Practicals in Physical Chemistry, A Modern Approach.

IPCH-3.4LAB: PHARMACEUTICAL ANALYSIS LAB
(Credits: 2, 3hrs/week)

All the experiments are to be repeated for at least three times and data generated for every experiment has to be entered in Microsoft excel file.

1. Determination of % purity of sodium benzoate.
2. Determination of λmax of $KMnO_4$
3. Determination of λmax of paracetamol
4. Preparation and standardization of iodine solution
5. Assay ferrous sulphate
6. Assay of norfloxacin tablets and chlorpromazine tablets
7. Experiment related to sampling of drugs from formulations
8. Experiment related to sampling of drugs from biological fluids
9. TLC profile of paracetamol and para aminophenol using various mobile phases
10. TLC profile of aspirin and salicylic acid using various mobile phases
11. Demonstration of column chromatography
12. Estimation of paracetamol using calibration curve method
13. Assay of paracetamol tablets and syrup
14. Estimation of effect of solvent, concentration on absorption of compounds
15. Demonstration of bathochromic shift and hypsochromic shift
16. Comparison of methods of assay of aspirin using titrimetry and spectroscopy

Reference Books

1. Gary D. Christian: Analytical Chemistry, 6th edition, John Wiley & sons Inc.
2. A. H. Becket and J. B. Stenlake: Practical Pharmaceutical Chemistry, 4th edn, the Athlone press
3. Henry P. Talbot: An introductory Course of Quantitative Chemical Analysis with Explanatory Notes, 6th edition.
4. G. H. Jeffery, J. Bassett, J. Mendhan, R. C. Denny: Vogel's text book of Quantitative Chemical Analysis.

IPCH-3.5LAB: CELL AND MOLECULAR BIOLOGY LAB
(Credits: 2, 3hrs/week)

1. Isolation of Mitochondria
2. Dialysis
3. Study of Mitosis (and microscopic observation of mitotic phases)
4. Study of Meiosis
5. Human karyotype
6. Identification of chromosomal abnormalities from karyotype
7. Problems on Reassociation Kinetics
8. Primer designing for PCR
9. Amplification of genes by PCR

IPCH-4.2T: CHEMISTRY OF NUTRACEUTICALS AND DAIRY PRODUCTS
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Nutrients

Food as a source of nutrients, functions of food, definition of nutrients, food pyramid, energy yielding and protective nutrients

Carbohydrates: Sources, functions, storage in the body and utilization of carbohydrates

Proteins: Sources, functions, essential and non-essential amino acids

Lipids: Classification, sources and requirements, essential fatty acids, functions of triglycerides, phospholipids, lipoproteins and cholesterol

Vitamins: Classification, sources and requirements, effects of deficiency of vitamins

Minerals: Sources, functions, bioavailability, effects of deficiency of minerals

Unit-II: Nutraceuticals

Introduction to nutraceuticals, definitions, basis of claims for a compound as a nutraceutical, regulatory issues for nutraceuticals including CODEX; concept of angiogenesis and the role of nutraceuticals; nutraceuticals for management of various diseases; nutraceuticals and their mechanisms of action and dosage levels.

Unit-III: Manufacture of Nutraceuticals

Manufacturing aspects of selected nutraceuticals such as lycopene, isoflavonoids, prebiotics and probiotics, glucosamine, phytosterols; formulation of functional foods containing nutraceuticals-stability issues, analytical issues and labelling issues; clinical testing of nutraceuticals and health foods; adverse effects and toxicity of nutraceuticals.

Unit-IV: Analysis of Milk and Milk Products

Sampling of milk and milk products, Roese-Gottlich fat analysis, determination of specific gravity, total solid, acidity, sediment test, analysis of butter-modified Kohman test, detection of added water, total solid-Marshall rennet test, methylene blue test, Reazurin test, alcohol test, Arnold Guaiac test, Phosphatase test.

Unit-V: Production of Milk Products

Manufacture of condensed milk, milk powder, cheese, ice-cream, butter, ghee, malted products, evaporated and dried products, their evaluation and quality parameters, defects encountered during production, packaging and storage.

Reference Books

1. Swaminathan, M: Advanced Text Book on Food and Nutrition, Vol-I, BAPPCO Publishers, Bangalore (1985).
2. Swaminathan, M: Advanced Text Book on Food and Nutrition, Vol-II, BAPPCO Publishers, Bangalore (1985).

3. Srilaxmi, B: Nutrition Science, New Age International (p) Ltd Publishers (2003).
4. Geoffrey, P: Dietary Supplements and Functional Foods, Blackwell Publishing (2006).
5. Losso, J. N: Angi-angiogenic Functional and Medicinal Foods, CRC Press (2007).
6. Cupp, J and Tracy, T. S: Dietary Supplements: Toxicology and Clinical Pharmacology, Humana Press (2003).
7. Manson, P: Dietary Supplements (2nd Ed) Pharmaceutical Press (2001).
8. Campbell, J. E and Summers, J. L: Dietary Supplement Labeling Compliance (2004).
9. Shi, J: Functional Food Ingredients and Nutraceuticals: Processing Technologies, Taylor & Francis Publ. CRC Press (2007).
10. Goldberg, I: Functional Foods: Designer Foods, Pharma foods, Nutraceuticals Chapman & Hall (1999).
11. Robert, E. C: Handbook of Nutraceuticals and Functional Foods, Wildman (2006).
12. Neeser, J. R and German, B. J: Bioprocesses and Biotechnology for Functional Foods and Nutraceuticals, Marcel Dekker (2004).
13. Milk and Milk Products, 4th Edn, Tata McGraw Hill Publishing Company Ltd, New Delhi (1973).
14. Jenness, R and Patton, S: Principles of Dairy Chemistry, John Wiley & Sons, New York
15. Webb, B. H; Johnson A. H and Alford J. A: Fundamentals of Dairy Chemistry, 2nd Edn, AVI, Westport, USA (1974).
16. Wong, N. P; Jenness, R; Kennedy, M and Marth, E. H: Fundamentals of Dairy Chemistry, 3rd Edn, Van Norstrand Reinhold, New York.

IPCH-4.3T: COORDINATION CHEMISTRY, CHEMICAL KINETICS, DILUTE SOLUTIONS & COLLIGATIVE PROPERTIES, COLLOIDS & SURFACE CHEMISTRY
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Transition Elements (3d series)

General group trends with special reference to electronic configuration, variable valency, colour, magnetic and catalytic properties, ability to form complexes and stability of various oxidation states (Latimer diagrams) for Mn, Fe and Cu.

Lanthanoids and Actinoids: Electronic configurations, oxidation states, colour, magnetic properties, lanthanide contraction, separation of lanthanides (ion exchange method only).

Unit-II: Coordination Chemistry

Valence Bond Theory (VBT): IUPAC system of nomenclature, Inner and outer orbital complexes of Cr, Fe, Co, Ni and Cu (coordination numbers 4 and 6). Structural and stereoisomerism in complexes with coordination numbers 4 and 6. Drawbacks of VBT

Crystal Field Theory

Crystal field effect, splitting of d-orbital in octahedral geometry (regular, compressed and elongated), Crystal field stabilization energy (CFSE), Crystal field effects for weak and strong fields, Tetrahedral symmetry, Factors affecting the magnitude of D (10 DQ), Spectrochemical series, Comparison of CFSE for Oh and Td complexes, Tetragonal distortion of octahedral geometry. Jahn-Teller distortion, square planar coordination, d-orbital splitting in square pyramidal, trigonal bipyramidal and linear geometries; applications of CFT

Unit-III: Chemical Kinetics

Introduction to chemical kinetics, rate of reaction, variation of concentration with time, rate laws and rate constant. Specific reaction rate, Factors influencing reaction rates: effect of concentration of reactants, effect of temperature, effect of pressure, effect of surface area of reactants, effect of radiation, effect of catalyst with simple examples, order of reaction. First order reaction, derivation of equation for rate constant. Characteristics of first order reaction. Units for rate constant. Half- life period, graph of first order reaction examples-decomposition of H_2O_2. Pseudo first order reaction examples-hydrolysis of methyl acetate, inversion of cane sugar, problems. Second order reaction, derivation of expression for 2^{nd} order rate constant, examples-saponification of ester. Characteristics of second order reaction, units for rate constants, half- life period and second order plots. Zero order reaction: derivation of rate expression, examples i)combination of H_2 and Cl_2 to form HCl, ii) thermal decomposition of HI on gold surface characteristics of Zero order reaction units of k, half-life period and graph, problems.

Methods of determination of order of reaction: i) method of integration, ii) half life method, iii) van't-Hoff differential method iv) Ostwald's isolation method. Problems

Effect of temperature on reaction rate: Arrhenius equation, temperature coefficient, concept of energy of activation, determination of energy of activation from Arrhenius equation and by graphical method,

problems. Simple collision theory based on hard sphere model, explanation of frequency factor, orientation or steric factor.

Unit-IV: Dilute Solutions and Colligative Properties

Dilute Solutions, Colligative Properties, Raoult's law, relative lowering of vapour pressure, molecular weight determination. Osmosis - laws of osmotic pressure, its measurement, determination of molecular weight from osmotic pressure. Elevation of boiling point and depression of freezing point. Derivation of relation between molecular weight and elevation in boiling point and depression in freezing point. Experimental methods for determining various colligative properties. Abnormal molar mass, Van't hoff factor, degree of dissociation and assocoation of solutes.

Unit-V: Colloids and Surface Chemistry

Definition of colloids. Classification of colloids. Solids in liquids (sols): preparations and properties-(including Kinetic, Optical and Electrical stability of colloids) Protective action. Hardy-Schultz law, Gold number. Liquids in liquids (emulsions): Types of emulsions, preparation and emulsifier. Liquids in solids (gels); Classification, preparations and properties, General applications of colloids. Micelles: Classification of surface active agents. Surfactant action, micellization and micellar interactions, Structure of micelles – spherical and laminar. Critical micellar concentration (CMC). Factors affecting the CMC of surfactants. Counter ion binding to micelles.

Adsorption: Types of adsorption, Factors influencing adsorption. Freundlich adsorption isotherm. Langmuir theory of unilayer adsorption isotherm. Applications**.**

Reference Books

1. Kotz, J. C., Treichel, P. M. & Townsend, J. R: General Chemistry Cengage Learning India Pvt. Ltd., New Delhi (2009).
2. Mahan, B. H: University Chemistry, 3rd Ed. Narosa (1998).
3. Petrucci, R. H: General Chemistry, 5th Ed. Macmillan Publishing Co.: New York (1985).
4. Cotton, F. A. & Wilkinson, G: Basic Inorganic Chemistry, Wiley.
5. Shriver, D. F. & Atkins, P. W: Inorganic Chemistry, Oxford University Press.
6. Wulfsberg, G: Inorganic Chemistry, Viva Books Pvt. Ltd.
7. Rodgers, G. E: Inorganic & Solid State Chemistry, Cengage Learning India Ltd (2008).
8. G. M. Barrow: Physical Chemistry Tata McGraw-Hill (2007).
9. G. W. Castellan: Physical Chemistry 4th Edn. Narosa (2004).
10. Prutton and Marron: Principles of Physical Chemistry.
11. Puri and Sharma: Text Book of Physical Chemistry.
12. S. K. Dogra: Physical Chemistry through Problems.
13. Lewis and Glasstone: Elements of Physical Chemistry.
14. Atkins: Physical Chemistry.
15. Puri, Sharma and Pathania: Text Book of Physical Chemistry.
16. Gurdeep Raj: Advanced Physical Chemistry.

IPCH-4.4T: PHARMACOLOY-1
(Units: 5, Hours: 60, Credits: 4)

Pharmacology-1 (Basics of Pharmacokinetics and Pharmacodynamics)

Note: The following units are essential to a pharmaceutical chemist to understand and design new molecules.

Unit-I: Drug Action at Receptors

Structure and classification of receptors, general mode of operation, Superfamily Type 1, 2, 3, 4, Ligand–receptor relationships, The chemical nature of the binding of ligands to receptors, Neurotransmitters, signaling process, Ion channels and their control, Membrane-bound enzymes-activation/deactivation, conformational changes in receptor, Binding groups, Position of binding, Size and shape, ligand concentration–response curves (Agonist concentration–response relationships, Antagonist concentration–receptor relationships), Ligand-receptor theories (Clark's occupancy theory, rate theory, two-state model), Citalopram, an antagonist antidepressant –case study, β-Blockers

Unit-II: ADME

Scheme of fate of dosage form after its administration, definition and introduction to concept of absorption, distribution, biotransformation and elimination of drug, Introduction to bioavailability and various equivalence referring plasma time profile of drug, significance of metabolisms involved in the absorption and bio transformation of drugs, effects of physico-chemical, pharmaceutical and biological factors on ADME, renal and non-renal excretion, Concept of clearance, disintegration and dissolution studies

Unit-III: Barriers to Drug Exposure in Living Systems

Introduction to Barriers, drug dosing, barriers in the mouth, stomach, gastrointestinal tract, kidney, permeation of the gastrointestinal cellular membrane, metabolism in the Intestine, enzymatic hydrolysis in the intestine, absorption enhancement in the Intestine, barriers in the blood stream, plasma enzyme hydrolysis, plasma protein binding, red blood cell binding, blood brain barrier,

Unit-IV: Plasma Membrane

The plasma membrane (Lipid, protein, carbohydrate components), Similarities and differences between plasma membranes in different cells, Cell walls, Bacterial cell and animal cell exterior surfaces, Virus, Tissue, Human skin, The transfer of species through cell membranes, (Osmosis, Filtration, Passive diffusion, Facilitated diffusion, Active transport, Endocytosis, Exocytosis), examples of drugs that affects the structure of cell membranes and walls (anti-infective agents, Local anaesthetics)

Unit-V: Chemotherapy

General principle of chemotherapy (various targets of chemotherapy covering pathology of infection and mechanism of actions of drugs, concept of resistance), Sulphonamides and co-trimoxazole,

Antibiotics-Penicillins, Cephalosporins, Chloramphenicol, Macrolides, Quinolines and Fluoroquinolins, Quinolones, Tetracyclines, Aminoglycosides and miscellaneous antibiotics; Chemotherapy of tuberculosis, leprosy, fungal diseases, viral diseases, AIDS, protozoal diseases, worm infections, urinary tract infections and sexually transmitted diseases.

Reference Books

1. Graham L. Patrick: An Introduction to Medicinal Chemistry, Oxford University Press.
2. Gareth Thomas: Medicinal Chemistry, 2nd edition, John Wiley & Sons Ltd.
3. Donald Cairns: Essentials of Pharmaceutical Chemistry, third edn, Pharmaceutical Press.
4. Edward H. Kerns and Li Di: Drug Like Properties: Concepts, Structure Design and Methods: from ADME to Toxicity Optimization, Elsevier Publications.
5. Bertram G. Katzung (International edition): Basic and Clinical pharmacology, Lange Medical Book / Mc Graw Hill, USA 2001 8th edition.
6. Rang H. P, Dale M.M and Ritter J.M, Churchill Livingston: Pharmacology, London, 4th edn
7. Goodman and Gilman's the Pharmacological Basis of Therapeutics (International edition) Mc Graw Hill, USA 2001 10th edition.
8. Braunwald, Fauci, Kasper, Hauser and Longo Jameson: Harrison's Principles of Internal Medicine (two volumes), 2001, Mc Graw Hill, New York, 15th edition.
9. H. Gerhard Vogel: Drug Discovery and Evaluation-Pharmacological Assays, 2nd edition, Springer Verlag, Berlin, Heidelberg.
10. A. H. Becket and J. B. Stenlake: Practical Pharmaceutical Chemistry, 4th edition, the Athlone press.
11. Sara E. Rosenbaum: Basic Pharmacokinetics and Pharmacodynamics: An Integrated Text Book and Computer Simulations, Wiley publications.

IPCH-4.5T: GENETIC ENGINEERING AND ITS APPLICATIONS IN PHARMACEUTICAL SCIENCES

(Units: 5, Hours: 60, Credits: 4)

Unit-I:

Concept of Genetic Engineering, Formulation of a recombinant protein production, Enzymes used in Gene Cloning-Restriction Endonucleases and their role in Gene Cloning, Nucleases, Ligases, Polymerases and DNA modifying enzymes; Characteristic features of Plasmids, Development of Vectors, Cloning and Expression vectors, Plasmid Vectors, Bacteriophage vectors, Cosmids, Phasmids, M-13 vectors Phagemids, Yeast vectors

Unit-II:

Construction of Genomic and c-DNA libraries, Labelling of Nucleic acids and antibodies (probes), Screening DNA Libraries for the desired clone- 1) Direct selection for the desired clone- Genetic selection 2) Identification of the clone from a gene library -Colony/ Plaque hybridization; Detection of Genes and Gene products- Southern Blotting, Northern Blotting, Western Blotting; Applications of Genetic Engineering- Production of recombinant hormones and vaccines by Genetically Engineered Microorganisms

Unit-III:

Laboratory requirements for Animal Cell Culture, Media designing, Types of media, Selection of media, Preparation, Sterilization of glassware, apparatus, reagents and media; Primary Cell Culture, Secondary Cell Culture; Disaggregation of tissues; Cell lines, Selection, maintenance and preservation of cell lines; Stem cell isolation, culture and application; Cell hybridization technology, Production of Monoclonal antibodies by Hybridoma Technology, application of Monoclonal antibodies in Pharmaceutical Sciences

Unit-IV:

Development of Transgenic Animals, Methods of DNA delivary-Microinjection, Calcium phosphate precipitation, Transformation with polyplexes, Transformation with liposomes and lipoplexes, Electroporation; Transgenic Mice technology, Molecular pharming-Production of Pharmaceutical products by Transgenic Animals, their (extraction and purification); Gene therapy: in-vivo and ex-vivo gene therapy with suitable examples

Unit-V:

Laboratory requirements for Plant Tissue Culture; MS media, Media sterilization, explant sterilization; Callus induction and its culture, Suspension culture; Micropropagation of pharmaceutically important plants; Production of Pharmaceutical products by Plant Tissue Culture, Development of Transgenic plants, Gene transfer to Plant Cells-Agrobacterium mediated transformation, Protoplast transformation, Particle bombardment; Transgenic Plants- Bt-Cotton, Transgenic Tomato-Flavr-Savor, Golden Rice

Reference Books

1. S. B. Primrose and R. M. Twyman: Principles of Gene Manipulation and Genomics.
2. T. A Brown: Gene Cloning and DNA Analysis.
3. Glick and Pasternak: Molecular Biotechnology: Principles and Applications.
4. Pratik Satya: Genomics and Genetic Engineering.
5. A. J Nair: Introduction to Biotechnology and Genetic Engineering.
6. Ian Freshney: Culture of Animal Cells.
7. Tom Strachen and A P Read: Human Molecular Genetics.
8. H. K. Das: Biotechnology.
9. S. S. Bhojwaninand and A. Razdan: Plant Tissue Culture

IPCH-4.3LAB: INORGANIC AND PHYSICAL CHEMISTRY LAB
(Credits: 2, 3hrs/week)

Inorganic Chemistry

Semi-micro qualitative analysis using H_2S of mixtures - not more than four ionic species (two anions and two cations and excluding insoluble salts) out of the following:

Cations: NH_4^+, Pb^{2+}, Ag^+ , Bi^{3+}, Cu^{2+}, Cd^{2+}, Sn^{2+}, Fe^{3+}, Al^{3+}, Co^{2+}, Cr^{3+}, Ni^{2+}, Mn^{2+}, Zn^{2+}, Ba^{2+}, Sr^{2+}, Ca^{2+}, K^+

Anions: CO_3^{2-} , S^{2-}, SO^{2-}, $S_2O_3^{2-}$, NO_3^-, CH_3COO^-, Cl^-, Br^-, I^-, NO_3^-,SO_4^{2-}, PO_4^{3-}, BO_3^{3-}, $C_2O_4^{2-}$, F^-

(Spot tests should be carried out wherever feasible)

1. Estimate the amount of nickel present in a given solution as bis(dimethylglyoximato) nickel(II) or aluminium as oximate in a given solution gravimetrically. 2. Draw calibration curve (absorbance at λ_{max} vs. concentration) for various concentrations of a given coloured compound ($KMnO_4$/$CuSO_4$) and estimate the concentration of the same in a given solution. 3. Determine the composition of the Fe^{3+}-salicylic acid complex solution by Job's method. 4. Estimation of (i) Mg^{2+} or (ii) Zn^{2+} by complexometric titrations using EDTA 5. Estimation of total hardness of a given sample of water by complexometric titration. 6. Determination of concentration of Na^+ and K^+ using Flame Photometry.

Physical Chemistry

Adsorption

1. Adsorption of acetic acid on animal charcoal, verification of Freundlich isotherm.

Colloidal Solution

1. Preparation of sol solution of arsenic sulphide and estimation of flocculation value for NaCl, KCl, $BaCl_2$, $AlCl_3$.

Chemical kinetics

1. Determination of rate of decomposition of hydrogen peroxide.
2. Determination of specific reaction rate of the hydrolysis of methyl acetate catalyzed by hydrogen ion at room temperature.
3. Comparison of strengths of HCl and H_2SO_4 in the kinetic study of hydrolysis of methyl acetate.
4. Comparative kinetics of Hydrolysis of methyl acetate and ethyl acetate in the presence of HCl.

Reference Books

1. Svehla, G. Vogel's Qualitative Inorganic Analysis, Pearson Education (2012).
2. Mendham, J. Vogel's Quantitative Chemical Analysis, Pearson (2009).
3. Khosla, B. D.; Garg, V. C. & Gulati, A: Senior Practical Physical Chemistry, R. Chand & Co.
4. Alexander Findlay: Practical Physical Chemistry
5. B. P. Levitt: Findlay's Practical Physical Chemistry
6. P. S. Sindhu: Practicals in Physical Chemistry, A Modern Approach

IPCH-4.4LAB: PHARMACOLOGY LAB
(Credits: 2, 3hrs/week)

1. Calculation of IC50 values and Ki values with given data using micro soft excel sheet (practice with at least five data sets)
2. Identification of agonist and antagonist using dose response curves (interpretation of dose response curves) (practice with at least 10 curves)
3. Determination of disintegration time of various types of tablet formulations of a drug
4. Determination of disintegration time of tablets of a drug from various manufacturers or batches (three experiments)
5. Drug concentration time profiles of various formulations at acidic and neutral buffers (at least 4 formulations)
6. Determination of minimum inhibitory concentration of antibacterial agents
7. Drawing of antibacterial agents using chem sketch programme and visualizing in 3D view

Reference Books

1. H. Gerhard Vogel: Drug Discovery and Evaluation-Pharmacological Assays, 2nd edition, Springer Verlag, Berlin, Heidelberg.
2. A. H. Becket and J. B. Stenlake: Practical Pharmaceutical Chemistry, 4th edition, the Athlone press
3. Sara E. Rosenbaum: Basic Pharmacokinetics and Pharmacodynamics: An integrated Text Book and Computer Simulations, Wiley publications

IPCH-4.5LAB: GENETIC ENGINEERING LAB
(Credits: 2, 3hrs/week)

1. Problems on Restriction Mapping
2. Restriction Digestion
3. Preparation of competent cells
4. Identification of competent cells by Blue-white screening
5. Preparation and sterilization of medium for animal cell culture
6. Culture and maintenance of cell lines
7. Viability tests for cell lines
8. Preparation and sterilization of MS Media for plant tissue culture
9. Induction of Callus
10. Suspension culture of plant cells
11. Production of Pharmaceutical products by plant tissue culture

Department of Pharmaceutical Chemistry, Telangana University
5yrs Integrated M.Sc Chemistry (Pharmaceutical Chemistry)
Syllabus for V and VI Semesters

The following syllabus for **5yrs Integrated M.Sc Chemistry (Pharmaceutical Chemistry)** was unanimously approved by board of studies in its meeting held at Department of Pharmaceutical Chemistry, Telangana University, Nizamabad on 13th July, 2017

Yr	Sem	Paper Code	Subject	hrs/ week	Marks		Duration of exam in hrs		Credits
					IE*	UE**	IE*	UE**	
III	**V**	**IPCH-5.1T**	Physical chemistry and topics of industrial importance	4	30	70	1	3	4
		IPCH-5.2T	Coordination and supramolecular chemistry	4	30	70	1	3	4
		IPCH-5.3T	Principles of drug discovery and development	4	30	70	1	3	4
		IPCH-5.4T	Molecular targets and signal transduction	4	30	70	1	3	4
		IPCH-5.5T	Principles of stereochemistry	4	30	70	1	3	4
		IPCH-5.3Lab	Principles of drug discovery and development lab	3	15	35	3	3	2
		IPCH-5.4Lab	Molecular targets and signal transduction lab	3	15	35	3	3	2
		IPCH-5.5Lab	Organic synthesis lab	3	15	35	3	3	2
				29	**195**	**455**			**26**

**University Examination; *Internal Examination

Members Present in the board of studies meeting held:

S.No	Name	Designation
1	Prof. Naseem, Dept. of Pharmaceutical Chemistry, Telangana University	Chairman
2	Dr. V. Chandrasekhar, Dept. of Pharmaceutical Chemistry, Telangana University	Ex-Officio Member
3	Dr. B. Shireesha, Dept. of Pharmaceutical Chemistry, Telangana University	Member
4	Dr. M. Satyanarayana, Dept. of Pharmaceutical Chemistry, Telangana University	Member
5	Prof. C. Veeresham, UCPSc, Kakatiya University	Member
6	Prof. G. Krishna Mohan, JNTU, Hyderabad	Member
7	Prof. V. Ravinder, Dept. of Chemistry, Kakatiya University	Member
8	Prof. A. Jayashree, JNTU, Hyderabad	Member
9	Prof. Faiz Ahmed Khan, IIT, Hyderabad	Member
10	Dr. S . Ramakrishna, Scientist-F, IICT, Hyderabad	Member
11	Dr. G. Kumaraswamy, Senior Principal Scientist, IICT, Hyderabad	Member
12	Prof. P. Veerasomaiah, Osmania University, Hyderabad	Member
13	Dr. A. Hari Padma Sree, Koti Women's College, Osmania University, Hyderabad	Member
14	Dr. B. Prabhasankar, Leads Pharma Pvt. Ltd, Hyderabad	Member
15	Dr. V. Jagan Mohan Reddy, Dr. Reddy's Laboratories Ltd, Hyderabad	Member
16	Dr. Abdul Rasheed Mohammed, Suven Life Sciences, Hyderabad	Member
17	Mr. N. Babu Rao, Telangana Foods, Hyderabad	Member
18	Mr. Tukaram, Student, Telangana University, Hyderabad	Member

IPCH-5.1T: PHYSICAL CHEMISTRY AND TOPICS OF INDUSTRIAL IMPORTANCE
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Thermodynamics

Thermodynamic relations-Gibbs equations, Maxwell relations, temperature dependence of G, Gibbs-Helmholtz equation; Pressure dependence of G, Chemical potential-Gibbs equation for non-equilibrium systems, Material equilibrium, Phase equilibrium, Clapeyron equation and Clausius-Clapeyron equation; Derivation of equilibrium constant from thermodynamic consideration, temperature dependence of equilibrium constant-the Van't Haff equation; Chemical potential of ideal gases, Solutions, Partial molar properties-significance; Relation between solution volume and partial molar volumes, Determination of partial molar volumes-slope and intercept methods, Variation of chemical potential with T and P; Gibbs-Duhem equation-derivation and significance, Ideal solutions and thermodynamic properties of ideal solution, Mixing quantities, Vapour pressure-Raoult's law, Thermodynamic properties of ideally dilute solutions, Vapour pressure-Henry's law, Non-ideal systems; Concept of fugacity, fugacity coefficient, determination of fugacity; Non ideal solution, Activities and activity coefficients, Standard-state conventions for non-ideal solutions, Excess functions and their determination, Determination of activity coefficients from vapour pressure measurements, Activity coefficients of nonvolatile solutes using Gibbs-Duhem equation.

Unit-II: Photochemistry-I

Interaction of radiation with matter, difference between thermal and photochemical processes; Law of Photochemistry, Grothus-Draper law and Stark-Einstein law, Quantum yield and its determination, examples of low and high quantum yields (Photochemical combination of hydrogen-bromine and hydrogen-chlorine). Jablonski diagram depicting various processes occurring in the excited state, qualitative description of fluorescence, phosphorescence, non- radioactive processes (Internal conversion, Intersystem crossing); Photosensitized reactions-energy transfer processes (simple examples); Sensitized fluorescence, Quenching of fluorescence, Stern-Volmer equation; Actinometry-ferrioxalate and uranyl oxalate actinometers

Unit-III: Surface Chemistry & Catalysis

BET theory of multilayer adsorption, derivation of BET equation, its limitations and its applications; Introduction to catalysis, characteristics of catalytic reactions, types of catalysis (homogeneous/heterogeneous/bio-/anchored/PTC), characteristics of catalyst, comparison of homogeneous and heterogeneous catalysis with examples, activation energy of catalyzed reactions; enzyme catalysis, characteristics of enzyme catalysis, classification of enzyme catalysis, kinetics of enzyme catalysis (Michaelis-Menten kinetics), factors effecting enzyme catalysis; steps in heterogeneous catalyzed reaction, mechanism of surface-catalyzed reactions, the Langmuir - Hinshelwood and the Eley-Rideal mechanism; Determination of surface area by BET method, Determination of pore size and pore volume distribution by BJH method; Some industrially important heterogeneous catalytic processes: heterogeneous catalytic hydrogenation and oxidation, Fischer-Tropsch synthesis, Water-Gas shift reactions, methanol economy.

Unit-IV: Nanomaterials

Overview of nanostructures and nanomaterials (natural and synthetic), definition and classification of nanodimensional materials/particles, important techniques for synthesis of nanomaterials (physical/chemical/biological, top-down/bottom-up); special focus on synthesis of gold and silver nanomaterials; self assembly/stabilization-Moore's theory; optical, electrical and magnetic properties;

characterization techniques (SEM, TEM, AFM, UV-Vis, XRD, POM); scope of applications of nanoparticles/nanomaterials in catalysis, energy, pharmacy, water treatment, food science, optoelectronics, fuel-cell and environment; Carbon related nanostructures (Fullerenes, Graphene, CNTs) properties & applications

Unit-V: Green Chemistry

Introduction, 12 principles of 'Green Chemistry' with appropriate examples, atom economy and scope, green solvents and choice of solvents-water as a solvent for organic reactions, ionic liquids, supercritical fluids, PEG, deep eutectic solvents and solventfree reactions with appropriate examples; Catalysis and green chemistry, phase transfer catalysis in green chemistry

Examples of Green Synthesis/ Reactions and some real world cases:

- **i)** Green Synthesis of the adipic acid, poly lactic acid and biodiesel
- **ii)** Microwave assisted reactions in water: Hofmann Elimination
- **iii)** Microwave assisted reactions in organic synthesis: Diels-Alder reaction
- **iv)** Ultrasound assisted reactions: Sonochemical Simmons-Smith Reaction (Ultrasonic alternative to Iodine)

Reference Books

1. Advanced Physical Chemistry-Gurdeep Raj
2. Principles of Physical Chemistry-Puri, Sharma and Pathania
3. Physical Chemistry-Walter Moore
4. Physical Chemistry-D. A. Mc Quarrie and J.D. Simon
5. A text book of Physical Chemistry-Glasstone and Van Nostrand
6. A. Zangwill, Physics at Surfaces, Cambridge Univ. Press, 1988.
7. Catalytic Chemistry, Wiley, 1992-B. Gates
8. Physical Chemistry of Surfaces, Wiley, 1997-A. W. Adamson, A. P. Gast
9. Principles and Practice of Heterogeneous Catalysis, 1997- J. M. Thomas and W.J. Thomas
10. Surface Science: Foundations of Catalysis and Nanoscience, Wiley, 2002- K. W. Kolasinski
11. Heterogeneous Catalysis, New Age, 2008-D.K. Chakrabarty and B. Viswanathan
12. Introduction to Surface Chemistry and Catalysis, Wiley, 2010-G. A. Somorjai and Y. Li
13. Physical chemistry of surfaces, 1990-Arthur W. Adamson
14. Chemical kinetics and catalysis Wiley-Interscience, 2001-R. I. Masel
15. An introduction to chemisorption and catalysis by metals, 1985-R. P. H. Gasser
16. Introduction to Nanotechnology John Wiley & Sons- Poole, C. P. & Owens, F. J.
17. Fundamentals and Applications of Nanomaterials, Artech House-Z. Guo, L. Tan
18. Nano: the essentials, Tata McGraw Hill-T. Pradeep
19. Nanostructured Materials and Nanotechnology, edited by Hari Singh Nalwa, Academic Press
20. Introduction to Nanotechnology-Charles P. Poole Jr, F. J. Owens
21. *New Trends in Green Chemistry*-Ahluwalia, V. K. & Kidwai, M. R

24. *Green Chemistry-Theory and Practical*, Oxford University Press Anastas, P.T. & Warner, J. K

25. Green Chemistry Experiments:

26. *A monograph* I.K. International Publishing House Pvt Ltd. New Delhi-Sharma, R. K.; Sidhwani, I. T. & Chaudhari, M. K
27. An Introduction to Green Chemistry, Vishal Publishing-V. Kumar,
28. Green Chemistry in the Pharmaceutical Industry, Wiley VCH-P. J. Dunn, A. S. Wells, M. T. Williams

IPCH-5.2T: COORDINATION AND SUPRAMOLECULAR CHEMISTRY
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Metal-Ligand Bonding Theories and Coordination Equilibria

Theories of metal bonding-VBT, CFT and their limitations (review), experimental evidence for metal-ligand bond covalency, molecular orbital (MO) theory of selected octahedral and tetrahedral complexes, sigma (σ)- and pi(π)-bonding in complexes, Nephelauxetic effect, Solvation of metal ions, formation of binary metal complexes and their stability, step-wise and overall stability constants relation, factors influencing the stability constants: (i) metal ion effects (charge, size and Irving - William's order) (ii) Ligand effects (basicity, chelate, macro cyclic and cryptate effects), Jahn Teller effect on stability constants of metal complexes, symbiosis (chemical), HSAB concept, Methods used for the determination of stability constants; pH metric and spectrophotometric methods. Ternary metal complexes formation, step wise and simultaneous equilibria with appropriate examples

Unit-II: Reaction Mechanisms of Transition Metal Complexes

Concept of labile and inert complexes in terms of VBT and CFT, types of ligand substitution reactions in octahedral (SN^1 dissociative & SN^2 associative), acid and base hydrolysis reactions of octahedral complexes, factors influencing acid hydrolysis, base hydrolysis and conjugate base (SN^1CB) mechanism, evidence in favour of S_N1CB mechanism, annation reactions, Ligand substitution reactions of square planar metal complexes, trans effect-theories and application; Electron transfer reactions: classification, inner and outer-sphere electron transfer reactions mechanism and evidences for one electron transfer, Marcus-Hush Theory, some examples of two electron transfer reactions.

Unit-III: Electronic Spectra and Magnetic Properties of Transition Metal Complexes

Electronic Spectra-UV-Vis, colors, charge transfer and d-d transitions, intensities and origin of spectra, metal centered electronic spectra of transition metal complexes: microstates, terms of p^2 and d^2 configurations, L-S (Russel-Saunders) coupling, spin-orbital coupling parameters, term symbols and splitting of terms in free atoms, hole formalism, energy ordering of terms (Hund's rules), selection rules for electronic transitions, Racah parameters-Orgel diagrams; Types of magnetic behavior (dia-para, ferro, ferri and anti-ferro), temperature effect on magnetism, magnetic susceptibility determination by Guoy's method, spin-orbit coupling, quenching of orbital angular momentum on A, E and T ground terms, general applications of magnetic metal compounds in materials and pharmaceutical chemistry

Unit-IV: Molecular Clusters

Main-group clusters: General synthesis, geometric and electronic structure, three, four and higher connect clusters, the *closo*-, *nido*-, *arachno*-borane structural paradigm, Wade-Mingos electron counting rules (polyhedral skeletal electron pair theory)

Transition-metal clusters: General synthesis, factors affecting the formation of metal-metal bond, structure, low nuclearity metal-carbonyl clusters, high nuclearity metal-carbonyl clusters (Fe, Ru, Os

Co, Rh, Ir), isolobal analogy, compounds with M-M multiple bonds structure and bonding in $[Re_2Cl_8]^{2-}$, Application of molecular clusters in catalysis, materials, and biology.

Unit-V: Supramolecular Chemistry

Introduction, synonymous terminology in supramolecular chemistry, types of guests (cationic, anionic neutral), types of non-covalent interactions (hydrogen bonding, electrostatic, ring π-π and metal-π interactions, Van der Waals, hydrophobic/solvophobic forces and closed shell interactions); Types of hosts, stability of host-guest complex/binding constants; Cation binding, binding between metal cation and macrocycles, factors affecting the hole size macrocyclic ligand; Anion binding, properties of anions, design principle, macrocycles with secondary binding sites; Neutral species binding, hydrogen bond receptors, complementary hydrogen bonding receptors, secondary interactions, supplementing hydrogen bonding with other interactions, dendrimers synthesis and applications; Applications in pharmacy, phase transfer reagents, molecular sensors, switches and molecular machinery, catalysis and nanotechnology.

Reference Books

1. Basic Inorganic Chemistry, Wiley publishers-Cotton, F. A & Wilkinson, G.
2. Inorganic Chemistry Principles of Structure and Reactivity, 4th Edn., Pearson Education India, 2006-J. E. Huheey, E. A. Keiter, R. A. Keiter
3. Mechanisms of Inorganic Reaction, John Wiley & Sons-F. Basolo, R. G. Pearson
4. Physical Methods in Chemistry, Saunders College-R. S. Drago,
5. Ligand Field Theory and its Applications, Wiley, India-B. N. Figgis, M. A. Hitchman
6. Selected Topics in Inorganic Chemistry-R D Madan, G D Tuli & Wahid U Malik
7. Concise Inorganic Chemistry, 4th Edn, Wiley-India-J. D. Lee,
8. Inorganic Chemistry, Oxford University Press-Shriver, D. F. & Atkins, P. W
9. Inorganic Electronic Spectroscopy, 2nd Edn, Elsevier publishers-A. B. P. Lever
10. The Chemistry of Metal Cluster Complexes-D. F. Shriver, H. D. Kaerz and R. D. Adams
11. Supramolecular Chemistry-Concepts and Perspectives (Wiley-VCH, 1995)-J. M. Lehn
12. Supramolecular Chemistry (Oxford University Press, 1999)-P. D. Beer, P. A. Gale, D. K. Smith
13. Supramolecular Chemistry (Wiley, 2000)-J. W. Steed and J. L. Atwood

IPCH-5.3T: PRINCIPLES OF DRUG DISCOVERY AND DEVELOPMENT
(Units: 5, Hours: 60, Credits: 4)

Unit-1: Drug Discovery-An Overview

Over view on drug discovery and drug development scenario; Statistics of drugs available; Overview on top selling drugs; neglected diseases; Sources of leads for drug development; Over view on drugs (with examples) obtained from ethnopharmaceutical sources, Plant sources, Marine sources, Microorganisms and Animal sources; Compound collections and data bases, merits and demerits of few natural leads quinine and curcumin

Unit-2: Pharmacokinetics and Drug Discovery

Introduction to Pharmacokinetics, PK Parameters (Volume of Distribution, Area Under the Curve, Clearance, Half-life, Bioavailability), General classification of pharmacokinetic properties, Drug regimens, Drug concentration analysis and its therapeutic significance, The importance of pharmacokinetics in drug discovery, Pharmacokinetic models, Intravascular administration and distribution, Extravascular administration, dissolution and absorption, Single oral dose, The calculation of t_{max} and C_{max}, Repeated oral doses, use of pharmacokinetics in drug design, Effects of Plasma Protein Binding on PK Parameters, Tissue Uptake, Drug distribution and 'survival', Drug design for pharmacokinetic problems (Variation of substituents, Stereoelectronic modifications etc)

Unit-3: Pharmacodynamics and Drug Discovery

Chemical stability, Metabolic stability and metabolic reactions; stereochemical, biological, and environmental factors effecting metabolism; species and metabolism; enzymes and metabolism; secondary implications of metabolism (Inactive metabolites, toxic metabolites, metabolites with similar activity, metabolites with different activity); sites of action; Hydrophilic/hydrophobic balance; Drug dose levels; Metabolic blockers; Removal of susceptible metabolic groups; 'Sentry' drugs—synergism; 'Search and destroy' drugs; Self-destruct drugs; kinetics of metabolites; metabolism and drug design (Delivery systems).

Unit-4: Drug-like Properties

Leads and analogues, Bioavailability, Solubility, Structure and Stability, Drug-like Properties (Lipinski Rules, Veber Rules), Application of rules for compound assessment, examples for drug properties and discovery inefficiencies, Property Profiling in Discovery, general chemical properties of functional groups, Drug-like Property Optimization in Discovery through prodrugs and bioisosters.

Unit-5: Structural Modification in Drug Discovery

Binding role of hydroxyl, amino groups, aromatic rings, double bonds; Structure-activity relationships of synthetic analogues with variation of substituents, extension of the structure, chain extensions/contractions, degree of unsaturation, ring expansions/contractions, ring variations, simplification of the structure, changing size and shape, rigidification of the structure; introduction of amino group, halogens, hydroxyl, carboxylic groups, sulphonic acid, thiols, sulphides, elements of luck and inspiration, case study of oxaminoquine, SAR of sulfonamides, local anesthetics,

Reference Books

1. Essentials of pharmaceutical chemistry, Donald Cairns, Pharmaceutical Press, ISBN 978-0 85369-745-9

2. Drug like properties: concepts, structure design and methods, from ADME to toxicity optimization, by Edward H. Kerns and Li Di, Elsevier publications, ISBN: 978-0-1236-9520-8
3. An introduction to medicinal chemistry, Graham L. Patrick, Oxford University Press
4. Medicinal Chemistry, 2nd edition, Gareth Thomas, John Wiley and sons publishers, ISBN 978-0-470-02597-0 (HB), 978-0-470-02598-7 (PB)
5. The practice of medicinal chemistry, Camille Georges Wermuth, Elsevier Publications, ISBN 0-12-744481-5
6. Foye's principles of medicinal chemistry, Thomos L. Lemke, David A. Williams, Lippincott Williams and Wilkin publishers,
7. Transport processes in pharmaceutical systems, Gordon L. Amidon, Ping I. Lee, Elizabeth M. Topp, Marcel Dekker Inc, ISBN: 0-8247-6610-5
8. Goodman and Gilman's Manual of pharmacology and therapeutics, McGraw Hill publications
9. Basic principles of drug discovery and development, Benzamin E. Blass, Elsevier publications, ISBN: 978-0-12-411508-8

IPCH-5.4T: MOLECULAR TARGETS AND SIGNAL TRANSDUCTION
(Units: 5, Hours: 60, Credits: 4)

Brief introduction and role of following targets in pathophysiology and treatment of various diseases to enable student to understand drug action and to facilitate him/ her to think about strategies of new drug discovery

Unit-1: Receptors as Drug Targets

Fundamentals of receptor action; ***GPCRs:*** Acetyl choline receptors (Muscarinic receptors), Adenosine receptors, Adrenoceptors, Purinergic receptors, dopamine receptors, angiotensin receptors; ***Direct ligand gated ion channel receptors:*** $GABA_A$ receptors, Acetyl choline receptors (Nicotinic receptors), Glutamate receptors; ***Cytokine receptors:*** TNF-α receptors, Interleukin receptors; ***Receptors associated with tyrosine kinase:*** Insulin receptors; ***Nuclear receptors (steroid hormone receptors):*** progesterone receptors, Mineralocorticoid receptors; ***other nuclear receptors:*** Peroxisome proliferator-activated receptor (PPAR)

Unit-2: Enzymes as Drug Targets

Fundamentals of enzyme action; ***Oxidoreductases:*** MAOs, COXs, LOXs, Aromatase, Dihydro folate reductase (DHFR), Xanthine Oxidase; ***Transferases:*** Protein Kinase-C, COMT, Reverse transcriptase, RNA Polymerase, DNA polymerase, Tyrosine kinases; ***Hydrolases (proteases):*** Aspartyl proteases; ***Hydrolases (metalloproteases):*** ACE, Human carboxy peptidase; ***Other hydrolases:*** Esterases (AChE, PDE, HDAC); ***Lyases:*** DOPA decarboxylase, carbonic anhydrase; Isomerases, DNA gyrase, topoisomerases; ***Ligases (Synthases):*** Thymidylate synthase

Unit-3: Ion Channels as Drug Targets:

Introduction; Voltage-gated Ca^{2+} channels (L-Type channels, T-Type channels); K^+ channels (Epithelial K^+ Channels, Voltage gated K^+ Channels); Na^+ channels (Epithelial Na^+channels (ENaC), Voltage-gated Na^+channels); Direct ligand gated ion channels; Ryanodine-inositol 1,4,5-triphosphate receptor Ca^{2+} channel (RIR-CaC) family: Rynodyne receptors, Cl^- channels, Acid sensing (Proton gated) Ion channels, Cyclic nucleotide gated (CNG) and hyperpolarization activated cyclic nucleotide gated (HCN) Channels, Transient receptor potential channels (Transient receptor potential Ca2+ channel (TRP-CC) family)

Unit-4: Membrane Transporters as Drug Targets, Various Physicochemical Mechanisms

Membrane transporters as drug targets: Fundamentals, effects of transporters, definition of terms: Uniporters, symporters, antiporters; Efflux transporters (MDR, ABC), Uptake transporters, Biogenic amine transporters, Excitatory amino acid transporters, GABA transporters, Glycine transporters, Na^+/H^+ antiporters, Proton pump, Cation-chloride co-transporter (CCC) family, Na^+/K^+ ATPase; drug action by various physicochemical mechanisms

Unit-5: Nucleic Acids, Ribosomes, Aquaporins as Drug Targets, Targets of Monoclonal Antibodies, Multi targeting Action of Drugs

Nucleic acids: DNA and RNA, Spindle as drug targets; ***Ribosome:*** 30S, 50S units as drug targets; ***Targets of monoclonal antibodies:*** TNF, Immunoglobulins, VEGF, EGFR; Aquaporins as drug targets; ***Multi targeting action of drugs:*** Introduction, Multi target pharmacology in Alzheimer's disease and cancer, basic considerations, its possibilities and limitations,

Reference Books

1. The handbook of receptor classification and signal transduction, Published by Sigma-Aldrich
2. Drug like properties: concepts, structure design and methods, from ADME to toxicity optimization, by Edward H. Kerns and Li Di, Elsevier publications, ISBN: 978-0-1236-9520-8
3. An introduction to medicinal chemistry, Graham L. Patrick, Oxford University Press
4. Drugs, their targets and the nature and number of drug targets, Peter Imming, Christian Sinning and Achim Meyer, Nature Reviews Drug Discovery, 5, 821-834, 2006, doi:10.1038/nrd2132
5. Drugs and their molecular targets: An updated overview, Yves Landry, Jean-Pierre Gies, Fundamental & Clinical Pharmacology 22 (2008) 1–18; doi:10.1111/j.1472-8206.2007.00548.x
6. Basic principles of drug discovery and development, Benzamin E. Blass, Elsevier publications, ISBN: 978-0-12-411508-8
7. www.animalsimulator.com
8. https://www.rcsb.org/

IPCH-5.5T: PRINCIPLES OF STEREOCHEMISTRY
(Units: 5, Hours: 60, Credits: 4)

Unit-I: Introduction

Isomerism definition, isomerism classification-constitutional isomerism and stereo isomerism; stereoisomerism-conformational isomerism and configurational isomerism; configurational isomerism-geometrical isomerism and optical isomerism

Constitutional isomerism: Definition and classification.

Molecular representations, symmetry and chirality: Wedge, Fischer, Sawhorse and Newman representations and their inter conversion; symmetry operations, symmetry elements (C_n & S_n) and criteria for chirality.

Conformational isomerism: Definition, conformations, stability and energy profile diagrams of ethane, propane and n-butane; conformations, stability and energy profile diagram of cyclobutane, cyclopentane and cyclohexane.

Unit-II: Conformational Analysis (Acyclic Systems)

Concept of dynamic stereochemistry, conformational enantiomers and conformational diastereomers; study of conformations of 1,2-disubstitued ethane derivatives like 1,2-dichloro ethane, ethylene chlorohydrin, ethylene glycol, butane-2,3-diol and (2-(N,N-dimethylamino)-1,2-diphenyl ethanol; study of conformations of unsaturated acyclic compounds like propylene, 1-butene, acetaldehyde, propionaldehyde; Klyne-Prelog terminology for torsion angles; quantitative analysis of conformation-reactivity relationship by the Winstein-Holness equation and the Curtin-Hammett principle; introduction to use of physical and spectral methods in conformational analysis.

Unit-III: Conformational Analysis (Cyclic Systems)

Study of conformations of mono, di and poly substituted cyclohexanes, cyclohexene, cyclohexanone (2-alkyl & 3-alkyl ketone effect) 2-halocyclohexanones, cycloheptane, cyclooctane; conformational analysis of fused bicyclic systems like bicyclo[3,3,0]octane, bicyclo[4,3,0]nonane (hydrindanes) and decalins; conformational analysis of heterocycles like piperidine, N-methylpiperidine, tropane, tropine and pseudotropine; reactivity of cyclohexane with axial and equatorial substituents; stereochemistry of addition to the carbonyl group of a rigid cyclohexanone ring

Unit-IV: Configurational Isomerism

Geometrical isomerism: Cis/Trans isomerism, E&Z nomenclature; spectral and chemical methods of configuration determination of E&Z isomers, determination of configuration of aldoximes and ketoximes

Optical isomerism: Definition, enantiomers, diastereomers, relationship between enantiomers and diastereomers; examples of asymmetric molecules (Glyceraldehyde, Lactic acid, Alanine) and disymmetric molecules (trans-1,2-dichlorocyclopropane); molecules with constitutionally symmetrical chiral carbons (Tartaric acid); molecules with constitutionally unsymmetrical chiral carbons (2,3-dibromopentane); R&S nomenclature (CIP), specific optical rotation (SOR), measurement of optical rotation, factors affecting the optical rotation

Unit-V: Techniques

Racemisation and resolution techniques: Racemisation methods (by mixing, synthesis and thermal) and resolution by direct crystallization, diastereoisomer salt formation and chiral chromatography

Determination of absolute configuration: Determination of absolute configuration by chemical correlation methods and by Mosher's method

Axial chirality, planar chirality and helicity: Stereochemistry of allenes, spiranes, chiral biphenyls (Atropisomerism), ansa compounds, trans-cyclooctane and helically chiral compounds

Reference Books

1. Stereochemistry and mechanism through solved problems - P. S. Kalsi
2. Stereochemistry of organic compounds - D. Nasipuri
3. Stereochemistry of carbon compounds - Ernest L Eliel
4. Stereochemistry conformation and mechanism - P. S. Kalsi (9^{th} edition)
5. Advanced organic chemistry - Jerry March
6. Advanced organic chemistry - Francis A Carey & Richard Sundberg
7. Organic chemistry - Clayden, Greeves, Warren & Wothers

IPCH-5.3LAB: PRINCIPLES OF DRUG DISCOVERY AND DEVELOPMENT LAB
(Credits: 2, 3hrs/week)

1. Molecular representations of few natural products
2. Determination of partition coefficient of various drugs
3. Prediction of physicochemical properties of few compound libraries
4. Determination of pharmacokinetic parameters of paracetamol using *in vitro* dissolution experiment
5. Investigations on stability of drugs in gastric environment and comparison with its prodrugs
6. Prediction of drug likeliness of few compound libraries
7. Prediction of metabolic stability of few compound libraries
8. In vitro absorption studies of various formulations of drugs
9. In vitro evaluation of compounds for antioxidant activity
10. Structure and activities of various drugs

IPCH-5.4LAB: MOLECULAR TARGETS AND SIGNAL TRANSDUCTION LAB
(Credits: 2, 3hrs/week)

1. Potential drug databases
2. Potential protein databases
3. Various protein visualizing softwares
4. Visualizing protein structures in protein databank
5. Visualizing protein ligand interactions with co-crystallized structures
6. Evaluation of drugs using various enzyme assays (in vitro)
7. Simulations of animal experiments demonstrating role of various targets in activity of drugs

IPCH-5.5LAB: ORGANIC SYNTHESIS LAB
(Credits: 2, 3hrs/week)

Synthesize the following compounds, purify and recrystallize as per the procedures. Analyze the synthesized compounds by TLC; melting points/boiling points by comparing with the standard samples. Calculate the moles, equivalents and yields.

Nitro benzene, *p*-Bromo acetanilide, *m*-Nitro aniline, Aspirin, Acetanilide, Benzylideneaniline, 2-Naphthyl methyl ether (Neroline), Nitrobenzoic acid and *n*-Butyl acetate

Reference Books

1. Vogel, A. I., Tatchell, A. R., Furnis, B. S., Hannaford, A. J. & Smith, P. W. G: Textbook of Practical Organic Chemistry, Prentice-Hall, 5th edition (1996).
2. Mann, F. G. & Saunders, B. C: Practical Organic Chemistry Orient-Longman (1960).
3. Ahluwalia, V.K. & Aggarwal, R: Comprehensive Practical Organic Chemistry.